

Volcanic Cones

Name: _____

Volcanic cones can be classified by their shapes. Label the three different kinds of volcanic cones pictured below. Label the parts of the volcanoes.



WORD BANK

Shield
vent

Composite
lava flow

Cinder
cinders

magma

Earth Science If8755

David Laing

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web molarity is the number of moles of a substance per litre of solution also known as molar concentration. A capital M signifies solutions labelled with molar concentration. A 1.0 M solution contains 1 mole of solute per litre of solution. Molality is the number of solute moles per kilogram.

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web 1. Molarity = $\frac{\text{moles of solute}}{\text{liters of solution}}$ = $\frac{0.4500}{0.825} = 0.540$ M
2. Molality = $\frac{\text{moles of solute}}{\text{kg of solvent}}$ = $\frac{0.4500}{0.0129} = 3.50$ M
3. Mole fraction = $\frac{\text{moles of solute}}{\text{moles of solute} + \text{moles of solvent}}$ = $\frac{0.4500}{0.4500 + 0.0129} = 0.980$
4. Molar concentration = $\frac{\text{moles of solute}}{\text{liters of solution}}$ = $\frac{0.4500}{0.0129} = 35.0$ M
5. Molar mass = $\frac{\text{mass of solute}}{\text{moles of solute}}$ = $\frac{141.9579}{0.4500} = 314.8$ g/mol
6. Mass of solute = $\text{Molar mass} \times \text{moles of solute}$ = $314.8 \times 0.4500 = 141.9579$ g

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web Apr 7 2014 Molarity = $\frac{\text{moles of solute}}{\text{liters of solution}}$ = $\frac{8.422}{0.250} = 33.7$ M
Calculate molality = $\frac{\text{moles of solute}}{\text{kg of solvent}}$ = $\frac{8.422}{0.100} = 84.2$ M
A solution with a molality of 2.0 M requires 2.0 mol of solute per liter. So $2.0 \times 1.0 = 2.0$ mol of solute per liter.

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web Answer: The correct option is C. Explanation: The molarity of a mixture can be calculated using the following formula:
$$M_{\text{mix}} = \frac{M_1 V_1 + M_2 V_2 + M_3 V_3 + \dots}{V_1 + V_2 + V_3 + \dots}$$
 where M is molarity, V is volume, and the subscripts $1, 2, 3, \dots$ represent the individual components of the mixture.

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2. What is the molarity of a solution containing 325 g of NaCl dissolved in 750 ml of solution?
3. 140 g of KCl is dissolved in 600 ml of water. What is the molarity?
4. 724.4 g of ammonium phosphate in 4500 ml of alcohol.

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web Determine the molarity for each of the following solutions:
1. 0.444 mol of COCl_2 in 0.654 l of solution.
2. 98.0 g of phosphoric acid (H_3PO_4) in 1.00 l of solution.
3. 2074 g of calcium hydroxide (Ca(OH)_2) in 40.00 ml of solution.
4. 10.5 kg of Na_2CO_3 in 2.00 l of solution.

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