



# HYDROGEN IONS AND ACIDITY

## Section Summary

### Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

### Key Vocabulary

<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>	<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>	<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>
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### Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

### Key Figures

- Figure 1: pH scale
- Figure 2: Ionization of water

### Principles of Ionization

Water is a neutral molecule in which the hydrogen and oxygen atoms are covalently bonded together. It is the ionization of these atoms that gives water its acidic and basic properties.

Water ionizes to form a hydrogen ion and a hydroxide ion. The concentration of these ions is  $10^{-7} \text{ M}$  at  $25^\circ\text{C}$ . The concentration of these ions is  $10^{-7} \text{ M}$  at  $25^\circ\text{C}$ .

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### Acid-Base Properties

<ul style="list-style-type: none"><li>Acid</li><li>Strong acid</li><li>Weak acid</li><li>Base</li><li>Strong base</li><li>Weak base</li><li>pH</li><li>Hydroxide ion</li><li>Conjugate acid</li><li>Conjugate base</li></ul>	<ul style="list-style-type: none"><li>Acid</li><li>Strong acid</li><li>Weak acid</li><li>Base</li><li>Strong base</li><li>Weak base</li><li>pH</li><li>Hydroxide ion</li><li>Conjugate acid</li><li>Conjugate base</li></ul>	<ul style="list-style-type: none"><li>Acid</li><li>Strong acid</li><li>Weak acid</li><li>Base</li><li>Strong base</li><li>Weak base</li><li>pH</li><li>Hydroxide ion</li><li>Conjugate acid</li><li>Conjugate base</li></ul>
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Hydrogen Ions Hubert Thomas Stanley Britton,1943    The ABC of Hydrogen Ion Control La Motte Chemical Products Company,William Anthony Taylor,1928

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