

Multiple Choice

Select the alternative that best completes the statement or answers the question. What is the most common primary gas used?

1. How many moles of butane gas (propose a molecular at typical conditions, 25°C)?
a. 1
b. 2
c. 3
d. 4
2. At standard $\Delta_f H^\circ$ (298 K) = 285.8 kJ/mol, what is the standard enthalpy of combustion?
a. 285.8
b. 571.6
c. 857.4
d. 1143.2
3. The standard enthalpy change in the forward rxn:
a. 285.8 kJ/mol
b. 571.6 kJ/mol
c. 857.4 kJ/mol
d. 1143.2 kJ/mol
4. Fuel butane gas $\Delta_f H^\circ$ (298 K) = 126.1 kJ/mol. How many moles of CO_2 gas are produced per mole of butane gas combusted at 298 K at constant P?
a. 2 mol
b. 3.5 mol
c. 4 mol
d. 5 mol
5. What is the state of the water in the following reaction? $\text{C}_2\text{H}_5\text{OH}(\text{l}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 3\text{H}_2\text{O}(\text{g})$
a. liquid
b. solid
c. gas
d. aqueous
6. What is the heat of combustion of butane?
a. 285.8 kJ/mol
b. 571.6 kJ/mol
c. 857.4 kJ/mol
d. 1143.2 kJ/mol
7. A compound's empirical formula is CH_2 , at this ratio, what is its molar mass? (The molar mass of CH_2 is 14.03 g/mol.)
a. 14.03
b. 28.06
c. 42.09
d. 56.12
8. Fuel butane gas $\Delta_f H^\circ$ (298 K) = 126.1 kJ/mol. How many moles of CO_2 gas are produced per mole of butane gas combusted at 298 K at constant P?
a. 2 mol
b. 3.5 mol
c. 4 mol
d. 5 mol
9. At standard $\Delta_f H^\circ$ (298 K) = 285.8 kJ/mol, what is the standard enthalpy of combustion?
a. 285.8
b. 571.6
c. 857.4
d. 1143.2
10. What is the standard enthalpy of formation of CO_2 at 298 K? (The standard enthalpy of formation of CO_2 is -393.5 kJ/mol.)
a. 393.5
b. 787.0
c. 1180.5
d. 1574.0
11. The standard enthalpy of formation of CO_2 at 298 K is -393.5 kJ/mol. How many moles of CO_2 gas are produced per mole of butane gas combusted at 298 K at constant P?
a. 2 mol
b. 3.5 mol
c. 4 mol
d. 5 mol

Multiple Choice Stoichiometry Test

John Moore



Multiple Choice Stoichiometry Test:

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