

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa for weak acids and bases.
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa for weak acids and bases.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                 |                                      |                 |
|-----------------|--------------------------------------|-----------------|
| acid            | acid dissociation constant ( $K_a$ ) | acid ionization |
| acidic solution | acidic solution                      | acidic solution |
| acidic solution | acidic solution                      | acidic solution |

## Key Equations

- |   |  |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

## Part A Completion

Write the appropriate words in the blanks to complete the following sentences. Use the words in the box to complete the sentences.

1. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .
2. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is  $1.0 \times 10^{-10}$ .
3. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .
4. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is  $1.0 \times 10^{-10}$ .
5. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .
6. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is  $1.0 \times 10^{-10}$ .
7. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .
8. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is  $1.0 \times 10^{-10}$ .
9. The  $K_a$  of a weak acid is  $1.0 \times 10^{-4}$ . The  $K_b$  of its conjugate base is  $1.0 \times 10^{-10}$ .
10. The  $K_b$  of a weak base is  $1.0 \times 10^{-4}$ . The  $K_a$  of its conjugate acid is  $1.0 \times 10^{-10}$ .

1.  $K_a$
2.  $K_b$
3.  $K_a$
4.  $K_b$
5.  $K_a$
6.  $K_b$
7.  $K_a$
8.  $K_b$
9.  $K_a$
10.  $K_b$

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**Kevin T. Patton**



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