

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and pK_a and pK_b for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and pK_a and pK_b for weak acids and bases.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

acid dissociation

acid dissociation

acid dissociation constant

acid dissociation

acid dissociation

acid dissociation

acid dissociation

acid

Key Equations

$K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1)

$K_b = \frac{[OH^-][BH^+]}{[B]}$ (Equation 19.2)

Part A Completion

Write the chemical formula for each of the following acids, bases, and salts, showing the correct number of atoms of each element in the formula.

1. K_a of acetic acid, CH_3COOH , is 1.8×10^{-5} .

2. K_b of ammonia, NH_3 , is 1.8×10^{-5} .

3. K_a of hydrofluoric acid, HF , is 6.6×10^{-4} .

4. K_b of pyridine, C_5H_5N , is 1.7×10^{-9} .

5. K_a of formic acid, $HCOOH$, is 1.8×10^{-4} .

6. K_b of methylamine, CH_3NH_2 , is 4.4×10^{-4} .

7. K_a of nitrous acid, HNO_2 , is 4.5×10^{-4} .

8. K_b of ethylamine, $CH_3CH_2NH_2$, is 5.6×10^{-4} .

9. K_a of benzoic acid, C_6H_5COOH , is 6.3×10^{-5} .

10. K_b of dimethylamine, $(CH_3)_2NH$, is 5.6×10^{-4} .

11. K_a of lactic acid, $CH_3CH(OH)COOH$, is 1.4×10^{-4} .

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Table of Contents 19 2 Section Review Hydrogen Ions And Acidity

1. Understanding the eBook 19 2 Section Review Hydrogen Ions And Acidity
 - The Rise of Digital Reading 19 2 Section Review Hydrogen Ions And Acidity
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Section Review Hydrogen Ions And Acidity
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 2 Section Review Hydrogen Ions And Acidity
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Section Review Hydrogen Ions And Acidity
 - Personalized Recommendations
 - 19 2 Section Review Hydrogen Ions And Acidity User Reviews and Ratings
 - 19 2 Section Review Hydrogen Ions And Acidity and Bestseller Lists
5. Accessing 19 2 Section Review Hydrogen Ions And Acidity Free and Paid eBooks
 - 19 2 Section Review Hydrogen Ions And Acidity Public Domain eBooks

- 19 2 Section Review Hydrogen Ions And Acidity eBook Subscription Services
- 19 2 Section Review Hydrogen Ions And Acidity Budget-Friendly Options
- 6. Navigating 19 2 Section Review Hydrogen Ions And Acidity eBook Formats
 - ePub, PDF, MOBI, and More
 - 19 2 Section Review Hydrogen Ions And Acidity Compatibility with Devices
 - 19 2 Section Review Hydrogen Ions And Acidity Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 19 2 Section Review Hydrogen Ions And Acidity
 - Highlighting and Note-Taking 19 2 Section Review Hydrogen Ions And Acidity
 - Interactive Elements 19 2 Section Review Hydrogen Ions And Acidity
- 8. Staying Engaged with 19 2 Section Review Hydrogen Ions And Acidity
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 2 Section Review Hydrogen Ions And Acidity
- 9. Balancing eBooks and Physical Books 19 2 Section Review Hydrogen Ions And Acidity
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 2 Section Review Hydrogen Ions And Acidity
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 19 2 Section Review Hydrogen Ions And Acidity
 - Setting Reading Goals 19 2 Section Review Hydrogen Ions And Acidity
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 19 2 Section Review Hydrogen Ions And Acidity
 - Fact-Checking eBook Content of 19 2 Section Review Hydrogen Ions And Acidity
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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