

Chemistry Study Guide

Periodic Table of the Elements																							
H	He																	Ne					
Li	Be																	B	C	N	O	F	Ne
Na	Mg																	Al	Si	P	S	Cl	Ar
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr						
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe						
Cs	Ba		Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn						
Fr	Ra		Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Uut	Fl	Uup	Lv	Uus	Uuo						
Lanthanide Series		La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu							
Actinide Series		Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr							

🔍 Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles 🧪

💡 Matter:

Anything that has **mass** and **occupies space** 🧪🌍

Examples: air, gases, liquids, solids!

🔬 Atoms:

Atoms are the **smallest, indivisible particle** 🧪👉

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

📌 Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

Section 2nuclear Chemistry Study Guide Key

Alan S. Wingrove, Robert L. Caret



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