

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. autoionization | 5. K_a | 6. K_b |
| 7. ionization constant | 8. K_a | 9. K_b |

Key Equation

- | | |
|-----------------------|-----------------------|
| 1. $[H^+] = 10^{-pH}$ | 2. $pH = -\log [H^+]$ |
| 3. $pH = -\log [H^+]$ | 4. $[H^+] = 10^{-pH}$ |

Part A Completion

Write the equation and/or the numerical value(s) of the constants indicated that are associated with the reaction described. Write the chemical equation for the reaction.

1. Write the equation for the autoionization of water.
2. Write the equation for the autoionization of water.
3. Write the equation for the autoionization of water.
4. Write the equation for the autoionization of water.
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18. Write the equation for the autoionization of water.
19. Write the equation for the autoionization of water.
20. Write the equation for the autoionization of water.

1. K_a
2. K_b
3. K_a
4. K_b
5. K_a
6. K_b
7. K_a
8. K_b
9. K_a
10. K_b
11. K_a
12. K_b
13. K_a
14. K_b
15. K_a
16. K_b
17. K_a
18. K_b
19. K_a
20. K_b

19 2 Hydrogen Ions And Acidity Worksheet Answers



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Hydrogen Ions Hubert Thomas Stanley Britton,1943 **The ABC of Hydrogen Ion Control** La Motte Chemical Products Company,William Anthony Taylor,1928

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