

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the K_a and K_b of acids and bases and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|---------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. neutral solution |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the unknown quantity that are indicated in the problem. Be sure to include units whenever a unit is appropriate or required.

1. What substance is _____, a weak hydrogen ion donor?
2. _____ is the concentration of H^+ ions present in a solution and equal to _____.
3. _____ is the concentration of H^+ ions in a solution.
4. _____ is the concentration of H^+ ions in a solution.
5. _____ is the concentration of H^+ ions in a solution.
6. _____ is the concentration of H^+ ions in a solution.
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19. _____ is the concentration of H^+ ions in a solution.
20. _____ is the concentration of H^+ ions in a solution.

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