

Multiple Choice:

Selecting the alternative that constitutes the statement or answers the question. What is most commonly observed with water?

1. How many moles of hydrogen gas are in a cylinder at typical conditions, 1000 L?
 a. 1
 b. 10
 c. 100
 d. 1000

2. At temperature $T_1 = 100^\circ\text{C}$, $\Delta H_{\text{vap}} = 40.7 \text{ kJ/mol}$, what is the enthalpy change of vaporization?
 a. 40.7
 b. 40.7
 c. 40.7
 d. 40.7

3. The enthalpy of fusion is the lowest for:
 a. Water at 0°C
 b. Ice at 0°C
 c. Liquid water at 0°C
 d. Solid water at 0°C

4. First Bornemann $\Delta H_f = 10.5 \text{ kJ/mol}$, $\Delta H_{\text{vap}} = 40.7 \text{ kJ/mol}$, how many moles of water are produced from the combustion of 100 g of methane?
 a. 10.5
 b. 40.7
 c. 10.5
 d. 40.7

5. What is the state of the water in the atmosphere?
 a. Liquid water
 b. Solid water
 c. Gaseous water
 d. Solid water

6. What is the boiling point of water?
 a. 100°C
 b. 100°C
 c. 100°C
 d. 100°C

7. A substance's specific heat is $1.0 \text{ J/g}^\circ\text{C}$, at this rate, how much heat is required to raise the temperature of 100 g of water from 100°C to 101°C?
 a. 100
 b. 100
 c. 100
 d. 100

8. First Bornemann $\Delta H_f = 10.5 \text{ kJ/mol}$, $\Delta H_{\text{vap}} = 40.7 \text{ kJ/mol}$, how many moles of water are produced from the combustion of 100 g of methane?
 a. 10.5
 b. 40.7
 c. 10.5
 d. 40.7

9. At temperature $T_1 = 100^\circ\text{C}$, $\Delta H_{\text{vap}} = 40.7 \text{ kJ/mol}$, what is the enthalpy change of vaporization?
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 c. 40.7
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10. The enthalpy of fusion is the lowest for:
 a. Water at 0°C
 b. Ice at 0°C
 c. Liquid water at 0°C
 d. Solid water at 0°C

11. The enthalpy of fusion is the lowest for:
 a. Water at 0°C
 b. Ice at 0°C
 c. Liquid water at 0°C
 d. Solid water at 0°C

Multiple Choice Stoichiometry Test

Mary Millhollon, Richard H. Langley



Multiple Choice Stoichiometry Test:

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