

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. autoionization | 5. K_a | 6. K_b |
| 7. ionization constant | 8. K_a | 9. K_b |

Key Equation

- | | |
|-----------------------|-----------------------|
| 1. $[H^+] = 10^{-pH}$ | 2. $pH = -\log [H^+]$ |
| 3. $pH = -\log [H^+]$ | 4. $[H^+] = 10^{-pH}$ |

Part A Completion

Write the equation and/or the numerical value(s) of the K_a and K_b for the acids and bases listed in the table. Also indicate which species are the conjugate acid and base.

1. H_2SO_4 is a strong acid. Write the equation and/or the numerical value(s) of K_a and K_b for the acids and bases listed in the table.
2. H_2SO_4 is a strong acid. Write the equation and/or the numerical value(s) of K_a and K_b for the acids and bases listed in the table.
3. H_2SO_4 is a strong acid. Write the equation and/or the numerical value(s) of K_a and K_b for the acids and bases listed in the table.
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8. H_2SO_4 is a strong acid. Write the equation and/or the numerical value(s) of K_a and K_b for the acids and bases listed in the table.
9. H_2SO_4 is a strong acid. Write the equation and/or the numerical value(s) of K_a and K_b for the acids and bases listed in the table.
10. H_2SO_4 is a strong acid. Write the equation and/or the numerical value(s) of K_a and K_b for the acids and bases listed in the table.

1.	H_2SO_4
2.	H_2SO_4
3.	H_2SO_4
4.	H_2SO_4
5.	H_2SO_4
6.	H_2SO_4
7.	H_2SO_4
8.	H_2SO_4
9.	H_2SO_4
10.	H_2SO_4

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KJ Lindholm-Leary



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