

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa, given  $K_a$  or  $K_b$
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa, given  $pK_a$  or  $pK_b$
3. Calculate the percent ionization of weak acids

## Vocabulary

- acid dissociation
- acid ionization

- acid ionization constant ( $K_a$ )
- acid ionization

- base ionization
- base ionization constant ( $K_b$ )

## Key Equations

- 1.  $K_a = \frac{[H^+][A^-]}{[HA]}$  (Equation 19.1)

- 2.  $K_b = \frac{[HB^+][OH^-]}{[B]}$  (Equation 19.2)

## Part A Completion

Write the equilibrium constant for the acid ionization of hydrofluoric acid and water. Identify the products and reactants in the equilibrium.

Write the equilibrium constant for the base ionization of ammonia.

Write the equilibrium constant for the acid ionization of acetic acid.

Write the equilibrium constant for the base ionization of ammonia.

Write the equilibrium constant for the acid ionization of hydrofluoric acid.

Write the equilibrium constant for the base ionization of ammonia.

Write the equilibrium constant for the acid ionization of acetic acid.

Write the equilibrium constant for the base ionization of ammonia.

Write the equilibrium constant for the acid ionization of hydrofluoric acid.

Write the equilibrium constant for the base ionization of ammonia.

Write the equilibrium constant for the acid ionization of acetic acid.

Write the equilibrium constant for the base ionization of ammonia.

1.  $K_a = \frac{[H^+][F^-]}{[HF]}$
2.  $K_b = \frac{[NH_4^+][OH^-]}{[NH_3]}$
3.  $K_a = \frac{[H^+][CH_3COO^-]}{[CH_3COOH]}$
4.  $K_b = \frac{[NH_4^+][OH^-]}{[NH_3]}$
5.  $K_a = \frac{[H^+][F^-]}{[HF]}$
6.  $K_b = \frac{[NH_4^+][OH^-]}{[NH_3]}$
7.  $K_a = \frac{[H^+][CH_3COO^-]}{[CH_3COOH]}$
8.  $K_b = \frac{[NH_4^+][OH^-]}{[NH_3]}$
9.  $K_a = \frac{[H^+][F^-]}{[HF]}$
10.  $K_b = \frac{[NH_4^+][OH^-]}{[NH_3]}$
11.  $K_a = \frac{[H^+][CH_3COO^-]}{[CH_3COOH]}$
12.  $K_b = \frac{[NH_4^+][OH^-]}{[NH_3]}$

# 19 2 Hydrogen Ions And Acidity Section Review

**WJ Hussar**



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