

# PHYSICS GRADE 10

## CHAPTER ONE LINEAR MOTION

### Introduction

Study of motion is divided into two;

1. Kinematics
2. Dynamics

In kinematics forces causing motion are disregarded while dynamics deals with motion of objects and the forces causing them.

### I. Displacement

*Distance moved by a body in a specified direction is called displacement.* It is denoted by letter 's' and has both magnitude and direction. Distance is the movement from one point to another. The SI unit for displacement is the metre (m).

### II. Speed

*This is the distance covered per unit time.*

**Speed = distance covered / time taken.** Distance is a scalar quantity since it has magnitude only. The SI unit for speed is metres per second (m/s or ms<sup>-1</sup>)

**Average speed = total distance covered / total time taken**

Other units for speed used are Km/h.

Examples

1. A body covers a distance of 10m in 4 seconds. It rests for 10 seconds and finally covers a distance of 50m in 60 seconds. Calculate the average speed.

*Solution*

Total distance covered = 10 + 50 = 100m

Total time taken = 4 + 10 + 6 = 20 seconds

Therefore average speed = 100 / 20 = 5m/s

2. Calculate the distance in metres covered by a body moving with a uniform speed of 180 km/h in 30 seconds.

*Solution*

Distance covered = speed \* time

$$= 180 * 1000 / 60 * 60 = 50 \text{ m/s}$$

$$= 50 * 30$$

$$= 1,500 \text{ m}$$

3. Calculate the time in seconds taken a by body moving with a uniform speed of 360km/h to cover a distance of 3,000 km?

*Solution*

Speed: 360 km/h = 360 \* 1000 / 60 \* 60 = 100m/s

Time = distance / speed

$$3000 * 1000 / 100$$

$$= 30,000 \text{ seconds.}$$

### III. Velocity

*This is the change of displacement per unit time.* It is a vector quantity.

**Velocity = change in displacement / total time taken**

The SI units for velocity are m/s

Examples

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## **Table of Contents 2014 Exemplar Question Of Physics Grade 10**

1. Understanding the eBook 2014 Exemplar Question Of Physics Grade 10
  - The Rise of Digital Reading 2014 Exemplar Question Of Physics Grade 10
  - Advantages of eBooks Over Traditional Books
2. Identifying 2014 Exemplar Question Of Physics Grade 10
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 2014 Exemplar Question Of Physics Grade 10
  - User-Friendly Interface
4. Exploring eBook Recommendations from 2014 Exemplar Question Of Physics Grade 10
  - Personalized Recommendations
  - 2014 Exemplar Question Of Physics Grade 10 User Reviews and Ratings
  - 2014 Exemplar Question Of Physics Grade 10 and Bestseller Lists

5. Accessing 2014 Exemplar Question Of Physics Grade 10 Free and Paid eBooks
  - 2014 Exemplar Question Of Physics Grade 10 Public Domain eBooks
  - 2014 Exemplar Question Of Physics Grade 10 eBook Subscription Services
  - 2014 Exemplar Question Of Physics Grade 10 Budget-Friendly Options
6. Navigating 2014 Exemplar Question Of Physics Grade 10 eBook Formats
  - ePub, PDF, MOBI, and More
  - 2014 Exemplar Question Of Physics Grade 10 Compatibility with Devices
  - 2014 Exemplar Question Of Physics Grade 10 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 2014 Exemplar Question Of Physics Grade 10
  - Highlighting and Note-Taking 2014 Exemplar Question Of Physics Grade 10
  - Interactive Elements 2014 Exemplar Question Of Physics Grade 10
8. Staying Engaged with 2014 Exemplar Question Of Physics Grade 10
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 2014 Exemplar Question Of Physics Grade 10
9. Balancing eBooks and Physical Books 2014 Exemplar Question Of Physics Grade 10
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 2014 Exemplar Question Of Physics Grade 10
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine 2014 Exemplar Question Of Physics Grade 10
  - Setting Reading Goals 2014 Exemplar Question Of Physics Grade 10
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 2014 Exemplar Question Of Physics Grade 10
  - Fact-Checking eBook Content of 2014 Exemplar Question Of Physics Grade 10
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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section 2 chemical equations a chemical equation uses symbols to represent a chemical reaction and shows the relationship between the reactants and the products of a reaction in words carbon dioxide and water react to form glucose and oxygen gas  $\text{CO}_2$

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dicarbon hexahydride oxygen carbon dioxide water synthesis equation copper sulfur copper i sulfide decomposition equation copper i sulfide copper sulfur single displacement equation iron lead ii nitrate iron iii nitrate