

2ND SEMESTER FINAL EXAM REVIEW GUIDE

This sheet may be used as good review of terms and concepts, but it is not "all inclusive." Study class notes, textbook, textbook CD, Unit Packets, Unit Packet Objectives, labs and **this guide**.

UNIT 7B – HUMAN EVOLUTION (9 Questions) Ch. 17 p. 368-378 and Ch. 13 p. 258-259

(Recommendation: Read all of these pages!)

Evolution of primates

Evolution of humans and human ancestors

Hominids: *Australopithecus*, *Homo habilis*, *Homo erectus*, *Homo sapiens*

Characteristics of primates

Bipedalism, opposable thumb, evolution of brain, hair, jaws, face, chin, parental care

Multiregional hypothesis vs. "Out of Africa" hypothesis

Hardy-Weinberg law – problems and definition (3 questions)

UNIT 8 – CLASSIFICATION & DIVERSITY (20 Questions) Ch. 14, 15, 16, 17

3 Domain System – list & describe 3 domains

Classification system – 7 taxon system - KPCOFGS

How are organisms classified?

Kingdoms – characteristics of

Binomial nomenclature

Phylogenetic tree

Invertebrate body patterns – types of symmetry

Invertebrates – characteristics of **each phyla**

Invertebrates: body segments, appendages, jointed or non, # of body openings, # of legs

Vertebrates – characteristics of **each class**

Endothermic vs. ectothermic

Mammals – characteristics of

Vertebrate hearts – chambers

single vs. double loop circulation and lungs vs. gills

oviparous, viviparous, ovoviviparous

UNIT 9 – PLANTS (33 questions) Ch. 16, 28, 29

Origin of plants

Characteristics of plant kingdom

Gymnosperms

Angiosperms

Monocot vs. Dicot

Growing seasons of plants – annuals, perennials, biennials

Flower - parts and functions

Double fertilization of flowers – steps of

Fruits – examples of, characteristics of, plant part

Seeds

Pollination

Leaf – parts of

Greenhouse Effect / global warming

Vascular system-xylem and phloem

Growth of a plant – stems and roots

Root – parts of, how growth occurs

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web the carbon hydrogen chains in vulcanized rubber are held together by two sulfur atoms that form covalent bonds between the chains these covalent bonds are commonly called disulfide bridges explore other molecules that have such disulfide bridges present your findings to the class

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web because of the fact that polar covalent bonds have higher melting and boiling points why do dipole forces have an impact on the melting and boiling points of substances this is because the dipole forces are an added force of attraction that slows the molecules down making it take more energy to melt or boil them