

# Acids & Bases HOMEWORK

## Homework

1. Propose the correct ion search of the strong acid listed below. Propose the incorrect ion search of the strong base listed below.

- a. **HF**: **Hydrofluoric acid**
- b. **HI**: **Hydroiodic acid**
- c. **HNO<sub>3</sub>**: **nitric acid**
- d. **HClO<sub>4</sub>**: **perchloric acid**
- e. **HBrO<sub>3</sub>**: **hypobromous acid**
- f. **HClO<sub>4</sub>**: **chloric acid**
- g. **H<sub>2</sub>SO<sub>4</sub>**: **hydrochloric acid**
- h. **H<sub>2</sub>SO<sub>4</sub>**: **hydrobromic acid**
- i. **H<sub>2</sub>SO<sub>4</sub>**: **hydroiodic acid**
- j. **H<sub>2</sub>SO<sub>4</sub>**: **hydrochloric acid**

2. Of the acids and bases listed below, write the ones that completely dissociate in water.

Acid	Base	Mixture	Neutral	Acidic	Basic	Neutral	Basic
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3. "Write a balanced equation for the reaction":  $\text{Ba(OH)}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + \text{H}_2\text{O}$

4.  $\text{H}_2\text{SO}_4$  will react with water and lose  $\text{H}^+$  ions in the process.

a. "Write a balanced equation for this reaction":  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{HSO}_4^-$

b. "Would you expect a large or small equilibrium constant for this reaction?" **small** because  $\text{H}_2\text{O}$  is a weak acid

c. "Would you expect  $\text{HSO}_4^-$  to be present almost all or almost none?" **present** because  $\text{H}_2\text{O}$  is strong at  $\text{pH} > 7$

d. "In your reaction equation, is  $\text{H}_2\text{SO}_4$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base?" **Brønsted-Lowry acid** because it is a proton donor

e. "Identify the conjugate base of  $\text{HSO}_4^-$ ":  $\text{SO}_4^{2-}$

f. "Identify  $\text{HSO}_4^-$  as a conjugate, an anticonjugate, or neither": **conjugate** because it can accept a proton. **Neither** because it is a strong acid

5.  $\text{H}^+$  will react with water and lose a  $\text{H}^+$  ion in the process.

a. "Write a balanced equation for this reaction":  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{OH}^-$

b. "Would you expect a large or small equilibrium constant for this reaction?" **large** because  $\text{H}_2\text{O}$  is a weak base

c. "Would you expect  $\text{OH}^-$  to be present almost all or almost none?" **present** because  $\text{H}_2\text{O}$  is strong at  $\text{pH} < 7$

d. "In your reaction equation, is  $\text{H}^+$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base?" **Brønsted-Lowry acid** because it is a proton donor

e. "Identify the conjugate acid of  $\text{OH}^-$ ":  $\text{H}_2\text{O}$

f. "Identify  $\text{OH}^-$  as a conjugate, an anticonjugate, or neither": **neither**

6.  $\text{H}_2\text{SO}_4$  will react with water and lose a  $\text{H}^+$  ion in the process.

a. "Write a balanced equation for this reaction":  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{HSO}_4^-$

b. "Would you expect a large or small equilibrium constant for this reaction?" **small** because  $\text{H}_2\text{O}$  is a weak acid

c. "Would you expect  $\text{HSO}_4^-$  to be present almost all or almost none?" **present** because  $\text{H}_2\text{O}$  is strong at  $\text{pH} > 7$

d. "In your reaction equation, is  $\text{H}_2\text{SO}_4$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base?" **Brønsted-Lowry acid** because it is a proton donor

e. "Identify the conjugate base of  $\text{HSO}_4^-$ ":  $\text{SO}_4^{2-}$

f. "Identify  $\text{HSO}_4^-$  as a conjugate, an anticonjugate, or neither": **conjugate** because it can accept a proton. **Neither** because it is a strong acid

## 19 3 Acids And Bases Answer Key Chemistry

**Henry Watts**

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