

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HBr~~ **hydrobromic acid**
- ~~HCl~~ **hydrochloric acid**
- ~~HNO₃~~ **nitric acid**
- ~~HClO₄~~ **perchloric acid**
- ~~Sulfuric hydrosulfate~~ **H₂SO₄**
- ~~potassium hydroxide~~ **KOH**
- ~~calcium hydroxide~~ **Ca(OH)₂**
- ~~calcium hydroxide~~ **Ca(OH)₂**
- ~~potassium hydroxide~~ **KOH**

HClO₃ is NOT a strong acid & does not completely dissociate in water.

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.

HCl

~~HF~~

~~HNO₂~~

~~Mg(OH)₂~~

HNO₃

NaOH

~~NaOH~~

NaClO₄

3. What is the formula for hydronium? **H₃O⁺**

4. Cu_2CO_3 will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{Cu}_2\text{CO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{CO}_3 + \text{Cu}_2\text{CO}_3$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because Cu_2CO_3 is a weak acid.**
- Would you define Cu_2CO_3 as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is Cu_2CO_3 acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of Cu_2CO_3 : **Cu_2CO_3**
- Identify Cu_2CO_3 as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's.**

5. F^- will react with water and gain a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{F}^- + \text{H}_2\text{O} \rightarrow \text{HF} + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F^- is a weak base.**
- Would you define F^- as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H^+ ion.**
- In your written equation, is F^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of F^- : **HF**
- Identify F^- as polyprotic, amphoteric, or neither: **neither**

6. HPO_4^{2-} will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{HPO}_4^{2-} + \text{H}_2\text{O} \rightarrow \text{H}_2\text{PO}_4^- + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HPO_4^{2-} is a weak acid.**
- Would you define HPO_4^{2-} as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is HPO_4^{2-} acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of HPO_4^{2-} : **PO_4^{3-}**
- Identify HPO_4^{2-} as polyprotic, amphoteric, or neither: **amphoteric, because it has a negative charge (and will like to accept a H^+) and it has H's (and will like to lose one).**

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Henry Watts



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