

WW2 topic quiz

Page 1 of 3

EASIER

(SOME) HARDER

Question 1: What date did Britain declare war on Germany?

- A. 3rd September 1938
- B. ~~3rd September 1939~~
- C. 3rd September 2039

- A. 13th September 1939
- B. ~~3rd September 1939~~
- C. 1st September 1939

Question 2: What date did the war in Europe end?

- A. 8th May 1944
- B. ~~6th May 1945~~
- C. 8th May 1946

- A. ~~8th May 1945~~
- B. 9th May 1945
- C. 10th May 1945

Question 3: What year did the USA join the war?

- A. 1940
- B. ~~1941~~
- C. 1942

Question 4: Which one of these statements is correct?

- A. Winston Churchill was the chancellor of Germany
- B. ~~Winston Churchill was the Prime Minister of Britain~~
- C. Winston Churchill was in a boy band called "WW2 Direction"

- A. Winston Churchill was replaced by Neville Chamberlain
- B. ~~Neville Chamberlain was replaced by Winston Churchill~~
- C. Winston Churchill was in a boy band called "WW2 Direction"

Question 5: Which one of these statements is correct?

- A. Albert Hitler was the German leader
- B. Rudolph Hitler was the German leader
- C. ~~Adolf Hitler was the German leader~~

- A. Adolf Hitler came to power in 1933 and invaded Poland in 1941
- B. Rudolph Hitler came to power in 1933 and invaded Poland in 1939
- C. ~~Adolf Hitler came to power in 1933 and invaded Poland in 1939~~

Question 6: What was the proper army name for the bombs that burst into flames on landing?

- A. Insect bombs
- B. ~~Incendary bombs~~
- C. Firebomb Fritz

Question 7: Why was the V2 so deadly compared to the V1?

- A. ~~It moved faster than sound so people couldn't hear it coming.~~
- B. It was thinner so it was harder to see
- C. It was much more thinner

Question 8: Why were children evacuated?

- A. To give their parents some peace
- B. For a nice holiday
- C. ~~To keep them safe~~

Question 9: Why were lampposts and pavements painted with white stripes?

- A. So that people knew where to park in the dark
- B. ~~So that people could see where they were going in the dark.~~
- C. So that enemy zebras could hide

Question 10: What was the popular phrase that the air raid wardens used?

- A. Put that cat out
- B. ~~Put that light out~~
- C. Put your left leg out and your left leg in

Multiple Choice Quiz World War 2

Jicheng Xie



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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a byproduct of the process as iron ions the ion meta dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web purpose in this experiment you will observe a single replacement reaction you will also use stoichiometry to predict what the theoretical yield of product is and calculate a percent yield reaction iron copper ii sulfate copper iron ii sulfate pre lab questions write the balanced equation

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

lab 4 lab 4 stoichiometry of iron copper ii sulfate redox - Jan 08 2023

web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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ions

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away cooper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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web stoichiometry lab report the reaction of iron with copper ii sulfate period 5 in this experiment i observed the reaction of iron with copper ii sulfate i obtained the results to this experiment through adding heated CuSO_4 solution to the iron powder stirring the solution adding water then removing water adding acetone heating the acetone on low

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