

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and  $pK_a$  and  $pK_b$  for weak acids and bases.
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and  $pK_a$  and  $pK_b$  for weak acids and bases.
3. Calculate the  $pH$  of weak acid and base solutions.

## Vocabulary

- |                 |                                      |                 |
|-----------------|--------------------------------------|-----------------|
| acid            | acid dissociation constant ( $K_a$ ) | acid ionization |
| acidic solution | acidic solution                      | acidic solution |
| acidic solution | acidic solution                      | acidic solution |

## Key Equations

- |                                 |                                  |
|---------------------------------|----------------------------------|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ | $K_b = \frac{[OH^-][HB^+]}{[B]}$ |
| $pH = -\log[H^+]$               | $pOH = -\log[OH^-]$              |

## Part A Completion

Write the chemical formula for each of the following acids, bases, and salts. Identify the acid, base, and salt in each of the following reactions.

1.  $HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$

2.  $H_2SO_4(aq) + NaOH(aq) \rightarrow Na_2SO_4(aq) + H_2O(l)$

3.  $HNO_3(aq) + NaOH(aq) \rightarrow NaNO_3(aq) + H_2O(l)$

4.  $H_2CO_3(aq) + NaOH(aq) \rightarrow Na_2CO_3(aq) + H_2O(l)$

5.  $H_2SO_4(aq) + NaOH(aq) \rightarrow Na_2SO_4(aq) + H_2O(l)$

6.  $HNO_3(aq) + NaOH(aq) \rightarrow NaNO_3(aq) + H_2O(l)$

7.  $H_2CO_3(aq) + NaOH(aq) \rightarrow Na_2CO_3(aq) + H_2O(l)$

8.  $H_2SO_4(aq) + NaOH(aq) \rightarrow Na_2SO_4(aq) + H_2O(l)$

9.  $HNO_3(aq) + NaOH(aq) \rightarrow NaNO_3(aq) + H_2O(l)$

10.  $H_2CO_3(aq) + NaOH(aq) \rightarrow Na_2CO_3(aq) + H_2O(l)$

11.  $H_2SO_4(aq) + NaOH(aq) \rightarrow Na_2SO_4(aq) + H_2O(l)$

1.  $HCl(aq)$
2.  $H_2SO_4(aq)$
3.  $HNO_3(aq)$
4.  $H_2CO_3(aq)$
5.  $H_2SO_4(aq)$
6.  $HNO_3(aq)$
7.  $H_2CO_3(aq)$
8.  $H_2SO_4(aq)$
9.  $HNO_3(aq)$
10.  $H_2CO_3(aq)$
11.  $H_2SO_4(aq)$

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**Linda Swisher, Gary A. Thibodeau**



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