

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on  $K_a$  and  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                        |                             |                   |
|------------------------|-----------------------------|-------------------|
| 1. autoionization      | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization          | 5. $K_a$                    | 6. $K_b$          |
| 7. ionization constant | 8. $pH$                     |                   |

## Key Equation

- |                       |                       |
|-----------------------|-----------------------|
| 1. $[H^+] = 10^{-pH}$ | 2. $pH = -\log [H^+]$ |
| 3. $pH = -\log [H^+]$ | 4. $[H^+] = 10^{-pH}$ |

## Part A Completion

Write the equation and/or the numerical value(s) of the constant(s) that are indicated in the spaces provided. Be sure to include proper units when appropriate.

1. Water autoionizes to  $H^+$  and hydroxide ions. The concentration of  $H^+$  in the autoionization of water at 25°C is  $1.0 \times 10^{-7}$  mol/L.
2. The autoionization constant for water is  $1.0 \times 10^{-14}$  at 25°C.
3. The  $K_a$  of acetic acid is  $1.8 \times 10^{-5}$ .
4. The  $K_b$  of ammonia is  $1.8 \times 10^{-5}$ .
5. The  $pH$  of a 0.1 M solution of acetic acid is 2.87.
6. The  $pH$  of a 0.1 M solution of ammonia is 11.13.
7. The  $K_a$  of acetic acid is  $1.8 \times 10^{-5}$ .
8. The  $K_b$  of ammonia is  $1.8 \times 10^{-5}$ .
9. The  $pH$  of a 0.1 M solution of acetic acid is 2.87.
10. The  $pH$  of a 0.1 M solution of ammonia is 11.13.

1.  $K_a$  of acetic acid2.  $K_b$  of ammonia3.  $K_a$  of acetic acid4.  $K_b$  of ammonia5.  $K_a$  of acetic acid6.  $K_b$  of ammonia7.  $K_a$  of acetic acid8.  $K_b$  of ammonia9.  $K_a$  of acetic acid10.  $K_b$  of ammonia11.  $K_a$  of acetic acid12.  $K_b$  of ammonia

# 19 2 Hydrogen Ions And Acidity Work Sheet Answers

**Bill Scott**



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**Hydrogen Ions** Hubert Thomas Stanley Britton, 1943      *The ABC of Hydrogen Ion Control* La Motte Chemical Products Company, William Anthony Taylor, 1928

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