

Multiple Choice

Selecting the alternative that constitutes the statement or answers the question. What is most commonly observed with water?

1. How many moles of hydrogen gas are in a cylinder at typical conditions, 1000 L?
 a. 1
 b. 10
 c. 100
 d. 1000

2. At temperature T_1 , 1 mol of H_2O vapor has kinetic energy E_1 . At temperature T_2 , 1 mol of H_2O vapor has kinetic energy E_2 .
 a. $E_2 = E_1$
 b. $E_2 = 2E_1$
 c. $E_2 = 4E_1$
 d. $E_2 = 8E_1$

3. The enthalpy of vaporization of water is 40.7 kJ/mol. The enthalpy of fusion of water is 6.01 kJ/mol. The enthalpy of sublimation of water is 46.7 kJ/mol.
 a. 40.7 kJ/mol
 b. 46.7 kJ/mol
 c. 52.7 kJ/mol
 d. 58.7 kJ/mol

4. Pure benzene at 1 atm and 25°C has a vapor pressure of 0.125 atm. The vapor pressure of pure benzene at 1 atm and 25°C is 0.125 atm.
 a. 0.125 atm
 b. 0.25 atm
 c. 0.375 atm
 d. 0.50 atm

5. Water is the most abundant molecule in the atmosphere. The most abundant molecule in the atmosphere is water.
 a. water
 b. oxygen
 c. nitrogen
 d. carbon dioxide

6. Water is the most abundant molecule in the atmosphere.
 a. 1.00 g/mol
 b. 1.80 g/mol
 c. 2.00 g/mol
 d. 2.20 g/mol

7. A sample of pure water at 25°C has a pH of 7.0. The pH of pure water at 25°C is 7.0.
 a. 7.0
 b. 7.0
 c. 7.0
 d. 7.0

8. Pure benzene at 1 atm and 25°C has a vapor pressure of 0.125 atm. The vapor pressure of pure benzene at 1 atm and 25°C is 0.125 atm.
 a. 0.125 atm
 b. 0.25 atm
 c. 0.375 atm
 d. 0.50 atm

9. At temperature T_1 , 1 mol of H_2O vapor has kinetic energy E_1 . At temperature T_2 , 1 mol of H_2O vapor has kinetic energy E_2 .
 a. $E_2 = E_1$
 b. $E_2 = 2E_1$
 c. $E_2 = 4E_1$
 d. $E_2 = 8E_1$

10. The enthalpy of vaporization of water is 40.7 kJ/mol. The enthalpy of fusion of water is 6.01 kJ/mol. The enthalpy of sublimation of water is 46.7 kJ/mol.
 a. 40.7 kJ/mol
 b. 46.7 kJ/mol
 c. 52.7 kJ/mol
 d. 58.7 kJ/mol

11. The enthalpy of vaporization of water is 40.7 kJ/mol. The enthalpy of fusion of water is 6.01 kJ/mol. The enthalpy of sublimation of water is 46.7 kJ/mol.

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