

## 2

## MATTER AND CHANGE

## SECTION 2.1 PROPERTIES OF MATTER (pages 39–42)

*This section helps you distinguish extensive from intensive properties and identify substances by their properties. It teaches you how to differentiate the three states of matter. It also defines a physical property and lists examples of physical properties and physical changes.*

## ► Describing Matter (page 39)

1. The \_\_\_\_\_ mass \_\_\_\_\_ of an object is a measure of the amount of matter the object contains.

2. How does an extensive property differ from an intensive property?

An extensive property depends on the amount of matter, while an intensive property depends on the type of matter.

## ► Identifying Substances (page 40)

3. Matter that has a uniform and definite composition is called a \_\_\_\_\_ substance \_\_\_\_\_.

4. Is the following sentence true or false? All samples of a substance have different physical properties. \_\_\_\_\_ false \_\_\_\_\_

5. A physical property is a quality or condition of a substance that can be \_\_\_\_\_ observed \_\_\_\_\_ or \_\_\_\_\_ measured \_\_\_\_\_ without changing the substance's composition.

6. Circle the letter of the term that is NOT a physical property.

- a. hardness      c. boiling point  
b. color      **d. melting**

7. Look at Table 2.1 on page 40. What is the melting point of bromine? \_\_\_\_\_  $-7^{\circ}\text{C}$  \_\_\_\_\_

8. Look at Table 2.1 on page 40. Circle the letter of the substance that is a yellow solid and melts at  $115^{\circ}\text{C}$ .

- a. sulfur**  
b. chlorine  
c. gold  
d. copper

## Section 3 1 Chemistry Answers

**Edward Wight Washburn**



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**Journal of the Society of Chemical Industry** Society of Chemical Industry (Great Britain), 1898 Includes list of members 1882-1902 and proceedings of the annual meetings and various supplements.

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