

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-----------------|--------------------------------------|-----------------|
| acid | acid dissociation constant (K_a) | acid ionization |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

Part A Completion

Write the appropriate words in the blanks to complete the following sentences. Use the words in the box to complete the sentences.

1. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
2. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
3. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
4. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
5. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
6. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
7. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
8. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .

9. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
10. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
11. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
12. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
13. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
14. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
15. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
16. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .

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Linda Swisher, Gary A. Thibodeau



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Table of Contents 19 2 Hydrogen Ions And Acidity Section Review

1. Understanding the eBook 19 2 Hydrogen Ions And Acidity Section Review
 - The Rise of Digital Reading 19 2 Hydrogen Ions And Acidity Section Review
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Hydrogen Ions And Acidity Section Review
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 2 Hydrogen Ions And Acidity Section Review
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Hydrogen Ions And Acidity Section Review
 - Personalized Recommendations
 - 19 2 Hydrogen Ions And Acidity Section Review User Reviews and Ratings

- 19 2 Hydrogen Ions And Acidity Section Review and Bestseller Lists
- 5. Accessing 19 2 Hydrogen Ions And Acidity Section Review Free and Paid eBooks
 - 19 2 Hydrogen Ions And Acidity Section Review Public Domain eBooks
 - 19 2 Hydrogen Ions And Acidity Section Review eBook Subscription Services
 - 19 2 Hydrogen Ions And Acidity Section Review Budget-Friendly Options
- 6. Navigating 19 2 Hydrogen Ions And Acidity Section Review eBook Formats
 - ePub, PDF, MOBI, and More
 - 19 2 Hydrogen Ions And Acidity Section Review Compatibility with Devices
 - 19 2 Hydrogen Ions And Acidity Section Review Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 19 2 Hydrogen Ions And Acidity Section Review
 - Highlighting and Note-Taking 19 2 Hydrogen Ions And Acidity Section Review
 - Interactive Elements 19 2 Hydrogen Ions And Acidity Section Review
- 8. Staying Engaged with 19 2 Hydrogen Ions And Acidity Section Review
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 2 Hydrogen Ions And Acidity Section Review
- 9. Balancing eBooks and Physical Books 19 2 Hydrogen Ions And Acidity Section Review
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 2 Hydrogen Ions And Acidity Section Review
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 19 2 Hydrogen Ions And Acidity Section Review
 - Setting Reading Goals 19 2 Hydrogen Ions And Acidity Section Review
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 19 2 Hydrogen Ions And Acidity Section Review
 - Fact-Checking eBook Content of 19 2 Hydrogen Ions And Acidity Section Review
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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