

The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

The total heat capacity of a molecular substance is the sum of each contribution (Figure 16.12 of the text). When equipartition is valid (when the temperature is well above the characteristic temperature of the mode  $T \gg \theta_m$ ) we can estimate the heat capacity by counting the numbers of modes that are active. In gases, all three translational modes are always active and contribute  $\frac{3}{2}R$  to the molar heat capacity. If we denote the number of active rotational modes by  $v_f^*$  (so for most molecules at normal temperatures  $v_f^* = 2$  for linear molecules, and 3 for non-linear molecules), then

# Physical Chemistry 9th Edition Instructors Solution Guide

**Jane E. Aaron, Robert A. Schwegler**

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Forthcoming Books Rose Arny, 2003 *An Introduction to Physical Science* James T. Shipman, Jerry D. Wilson, Aaron W. Todd, 2002-07-08 This text presents a survey of the physical sciences physics chemistry astronomy meteorology and geology for non science majors. Topics are treated both descriptively and quantitatively providing flexibility for instructors who wish to emphasize a highly descriptive approach a highly quantitative approach or any spectrum in between.

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