

# Chemistry Study Guide

# Periodic Table of the Elements

<small>1</small> <b>H</b> <small>1.008</small>																	<small>2</small> <b>He</b> <small>4.003</small>	
<small>3</small> <b>Li</b> <small>6.941</small>	<small>4</small> <b>Be</b> <small>9.012</small>											<small>5</small> <b>B</b> <small>10.81</small>	<small>6</small> <b>C</b> <small>12.01</small>	<small>7</small> <b>N</b> <small>14.01</small>	<small>8</small> <b>O</b> <small>16.00</small>	<small>9</small> <b>F</b> <small>18.998</small>	<small>10</small> <b>Ne</b> <small>20.18</small>	
<small>11</small> <b>Na</b> <small>22.99</small>	<small>12</small> <b>Mg</b> <small>24.31</small>											<small>13</small> <b>Al</b> <small>26.98</small>	<small>14</small> <b>Si</b> <small>28.09</small>	<small>15</small> <b>P</b> <small>30.97</small>	<small>16</small> <b>S</b> <small>32.07</small>	<small>17</small> <b>Cl</b> <small>35.45</small>	<small>18</small> <b>Ar</b> <small>39.95</small>	
<small>19</small> <b>K</b> <small>39.10</small>	<small>20</small> <b>Ca</b> <small>40.08</small>	<small>21</small> <b>Sc</b> <small>44.96</small>	<small>22</small> <b>Ti</b> <small>47.88</small>	<small>23</small> <b>V</b> <small>50.94</small>	<small>24</small> <b>Cr</b> <small>52.00</small>	<small>25</small> <b>Mn</b> <small>54.94</small>	<small>26</small> <b>Fe</b> <small>55.85</small>	<small>27</small> <b>Co</b> <small>58.93</small>	<small>28</small> <b>Ni</b> <small>58.71</small>	<small>29</small> <b>Cu</b> <small>63.55</small>	<small>30</small> <b>Zn</b> <small>65.39</small>	<small>31</small> <b>Ga</b> <small>69.72</small>	<small>32</small> <b>Ge</b> <small>72.64</small>	<small>33</small> <b>As</b> <small>74.92</small>	<small>34</small> <b>Se</b> <small>78.96</small>	<small>35</small> <b>Br</b> <small>79.90</small>	<small>36</small> <b>Kr</b> <small>83.80</small>	
<small>37</small> <b>Rb</b> <small>85.47</small>	<small>38</small> <b>Sr</b> <small>87.62</small>	<small>39</small> <b>Y</b> <small>88.91</small>	<small>40</small> <b>Zr</b> <small>91.22</small>	<small>41</small> <b>Nb</b> <small>92.91</small>	<small>42</small> <b>Mo</b> <small>95.94</small>	<small>43</small> <b>Tc</b> <small>98.91</small>	<small>44</small> <b>Ru</b> <small>101.07</small>	<small>45</small> <b>Rh</b> <small>102.91</small>	<small>46</small> <b>Pd</b> <small>106.42</small>	<small>47</small> <b>Ag</b> <small>107.87</small>	<small>48</small> <b>Cd</b> <small>112.41</small>	<small>49</small> <b>In</b> <small>114.82</small>	<small>50</small> <b>Sn</b> <small>118.71</small>	<small>51</small> <b>Sb</b> <small>121.76</small>	<small>52</small> <b>Te</b> <small>127.60</small>	<small>53</small> <b>I</b> <small>126.91</small>	<small>54</small> <b>Xe</b> <small>131.29</small>	
<small>55</small> <b>Cs</b> <small>132.91</small>	<small>56</small> <b>Ba</b> <small>137.33</small>			<small>72</small> <b>Hf</b> <small>178.49</small>	<small>73</small> <b>Ta</b> <small>180.95</small>	<small>74</small> <b>W</b> <small>183.84</small>	<small>75</small> <b>Re</b> <small>186.21</small>	<small>76</small> <b>Os</b> <small>190.23</small>	<small>77</small> <b>Ir</b> <small>192.22</small>	<small>78</small> <b>Pt</b> <small>195.08</small>	<small>79</small> <b>Au</b> <small>196.97</small>	<small>80</small> <b>Hg</b> <small>200.59</small>	<small>81</small> <b>Tl</b> <small>204.38</small>	<small>82</small> <b>Pb</b> <small>207.2</small>	<small>83</small> <b>Bi</b> <small>208.98</small>	<small>84</small> <b>Po</b> <small>209</small>	<small>85</small> <b>At</b> <small>210</small>	<small>86</small> <b>Rn</b> <small>222</small>
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## Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

## Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

## Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons ( $e^-$ )
- **Inner part** → *nucleus* → houses the protons ( $p^+$ ) and neutrons ( $n$ )
- Outer & Inner part → together these are the **subatomic particles**

## Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	$p^+$	1	+1	Nucleus
Neutron	$n$	1	0	Nucleus
Electron	$e^-$	$\sim 0$	-1	Orbit/Shell

# Section 2 nuclear Chemistry Study Guide Key

**Wolfgang Guggemos**



## Section 2nuclear Chemistry Study Guide Key:

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This book delves into Section 2nuclear Chemistry Study Guide Key. Section 2nuclear Chemistry Study Guide Key is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Section 2nuclear Chemistry Study Guide Key, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Section 2nuclear Chemistry Study Guide Key
- Chapter 2: Essential Elements of Section 2nuclear Chemistry Study Guide Key
- Chapter 3: Section 2nuclear Chemistry Study Guide Key in Everyday Life
- Chapter 4: Section 2nuclear Chemistry Study Guide Key in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Section 2nuclear Chemistry Study Guide Key. This chapter will explore what Section 2nuclear Chemistry Study Guide Key is, why Section 2nuclear Chemistry Study Guide Key is vital, and how to effectively learn about Section 2nuclear Chemistry Study Guide Key.
3. In chapter 2, the author will delve into the foundational concepts of Section 2nuclear Chemistry Study Guide Key. The second chapter will elucidate the essential principles that must be understood to grasp Section 2nuclear Chemistry Study Guide Key in its entirety.
4. In chapter 3, the author will examine the practical applications of Section 2nuclear Chemistry Study Guide Key in daily life. This chapter will showcase real-world examples of how Section 2nuclear Chemistry Study Guide Key can be effectively utilized in everyday scenarios.
5. In chapter 4, the author will scrutinize the relevance of Section 2nuclear Chemistry Study Guide Key in specific contexts. The fourth chapter will explore how Section 2nuclear Chemistry Study Guide Key is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, this book will draw a conclusion about Section 2nuclear Chemistry Study Guide Key. The final chapter will summarize the key points that have been discussed throughout the book.
- The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Section 2nuclear Chemistry Study Guide Key.

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