



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Define acid and base and relate them to the pH of a solution.
- Define strong and weak acids.

Key Vocabulary

<ul style="list-style-type: none">concentrationhydrogen ionacidbasestrong acidweak acid	<ul style="list-style-type: none">concentrationhydrogen ionacidbasestrong acidweak acid
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Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

Key Figures

- Figure 1: pH scale
- Figure 2: Ion concentration

Points of Consideration

What are the components of acids and bases? How do acids and bases affect the pH of a solution? How do acids and bases affect the solubility of other substances?

What is the difference between a strong acid and a weak acid? How do the concentrations of strong acids and weak acids differ?

What is the difference between a strong base and a weak base? How do the concentrations of strong bases and weak bases differ?

What is the difference between an acid and a base? How do the properties of acids and bases differ?

What is the difference between a strong acid and a weak acid?

Answers

1. $\text{pH} = -\log_{10}[\text{H}^+]$	2. $\text{pH} = \log_{10}[\text{OH}^-]$
3. $\text{pH} = 7$	4. $\text{pH} = 14$
5. $\text{pH} = 1$	6. $\text{pH} = 13$
7. $\text{pH} = 3$	8. $\text{pH} = 11$
9. $\text{pH} = 5$	10. $\text{pH} = 9$
11. $\text{pH} = 7$	12. $\text{pH} = 10$
13. $\text{pH} = 8$	14. $\text{pH} = 12$
15. $\text{pH} = 6$	16. $\text{pH} = 11$
17. $\text{pH} = 4$	18. $\text{pH} = 10$
19. $\text{pH} = 2$	20. $\text{pH} = 11$
21. $\text{pH} = 1$	22. $\text{pH} = 12$
23. $\text{pH} = 3$	24. $\text{pH} = 13$
25. $\text{pH} = 5$	26. $\text{pH} = 11$
27. $\text{pH} = 7$	28. $\text{pH} = 10$
29. $\text{pH} = 9$	30. $\text{pH} = 12$
31. $\text{pH} = 11$	32. $\text{pH} = 13$
33. $\text{pH} = 13$	34. $\text{pH} = 14$
35. $\text{pH} = 14$	36. $\text{pH} = 14$

19 2 Hydrogen Ions And Acidity Work Sheet Answers

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