



# HYDROGEN IONS AND ACIDITY

## Section Summary

### Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

### Key Vocabulary

<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>	<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>	<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li><li>base</li><li>strong base</li><li>weak base</li><li>pH</li><li>hydroxide ion</li><li>conjugate acid</li><li>conjugate base</li></ul>
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### Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

### Key Figures

- Figure 1: pH scale
- Figure 2: Ionization of water

### Principles of Ionization

Water is a neutral molecule in which the hydrogen and oxygen atoms are covalently bonded together. It is the ionization of these atoms that gives water its acidic and basic properties.

Water ionizes to form a hydrogen ion and a hydroxide ion. The concentration of these ions is  $10^{-7} \text{ M}$  at  $25^\circ\text{C}$ . The concentration of these ions is  $10^{-7} \text{ M}$  at  $25^\circ\text{C}$ .

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### Acid-Base Properties

<ul style="list-style-type: none"><li>Acid</li><li>Strong acid</li><li>Weak acid</li><li>Base</li><li>Strong base</li><li>Weak base</li><li>pH</li><li>Hydroxide ion</li><li>Conjugate acid</li><li>Conjugate base</li></ul>	<ul style="list-style-type: none"><li>Acid</li><li>Strong acid</li><li>Weak acid</li><li>Base</li><li>Strong base</li><li>Weak base</li><li>pH</li><li>Hydroxide ion</li><li>Conjugate acid</li><li>Conjugate base</li></ul>	<ul style="list-style-type: none"><li>Acid</li><li>Strong acid</li><li>Weak acid</li><li>Base</li><li>Strong base</li><li>Weak base</li><li>pH</li><li>Hydroxide ion</li><li>Conjugate acid</li><li>Conjugate base</li></ul>
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## 19 2 Hydrogen Ions And Acidity Worksheet Answers

**Michal Rosen-Zvi**

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**Hydrogen Ions** Hubert Thomas Stanley Britton,1943    **The ABC of Hydrogen Ion Control** La Motte Chemical Products Company,William Anthony Taylor,1928

## **19 2 Hydrogen Ions And Acidity Worksheet Answers** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has been more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**19 2 Hydrogen Ions And Acidity Worksheet Answers**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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