

# Acids & Bases HOMEWORK

## Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HBr~~ **hydrobromic acid**
- ~~HCl~~ **hydrochloric acid**
- ~~HNO<sub>3</sub>~~ **nitric acid**
- ~~HClO<sub>4</sub>~~ **perchloric acid**
- ~~Sulfuric hydrosulfate~~ **H<sub>2</sub>SO<sub>4</sub>**
- ~~potassium hydroxide~~ **KOH**
- ~~calcium hydroxide~~ **Ca(OH)<sub>2</sub>**
- ~~calcium hydroxide~~ **Ca(OH)<sub>2</sub>**
- ~~potassium hydroxide~~ **KOH**

~~HClO<sub>4</sub> is a STRONG acid & it does not completely dissociate in water.~~

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.

**HCl**

~~HF~~

~~HNO<sub>2</sub>~~

~~Mg(OH)<sub>2</sub>~~

**HNO<sub>3</sub>**

**NaOH**

~~NaOH~~

**NaClO<sub>4</sub>**

3. What is the formula for hydronium? **H<sub>3</sub>O<sup>+</sup>**

4.  $\text{Cu}_2\text{CO}_3$  will react with water and lose a  $\text{H}^+$  ion in the process.

- Write a balanced equation for this reaction:  $\text{Cu}_2\text{CO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{CO}_3 + \text{Cu}_2\text{CO}_3$
- Would you expect a large or small equilibrium constant for this reaction? **small, because  $\text{Cu}_2\text{CO}_3$  is a weak acid.**
- Would you define  $\text{Cu}_2\text{CO}_3$  as a proton donor or a proton acceptor? **proton donor, because it is losing a  $\text{H}^+$  ion.**
- In your written equation, is  $\text{Cu}_2\text{CO}_3$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of  $\text{Cu}_2\text{CO}_3$ :  **$\text{Cu}_2\text{CO}_3$**
- Identify  $\text{Cu}_2\text{CO}_3$  as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple hydrogens.**

5.  $\text{F}^-$  will react with water and gain a  $\text{H}^+$  ion in the process.

- Write a balanced equation for this reaction:  $\text{F}^- + \text{H}_2\text{O} \rightarrow \text{HF} + \text{OH}^-$
- Would you expect a large or small equilibrium constant for this reaction? **small, because  $\text{F}^-$  is a weak base.**
- Would you define  $\text{F}^-$  as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a  $\text{H}^+$  ion.**
- In your written equation, is  $\text{F}^-$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of  $\text{F}^-$ : **HF**
- Identify  $\text{F}^-$  as polyprotic, amphoteric, or neither. **neither**

6.  $\text{HSO}_4^-$  will react with water and lose a  $\text{H}^+$  ion in the process.

- Write a balanced equation for this reaction:  $\text{HSO}_4^- + \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_4 + \text{SO}_4^{2-}$
- Would you expect a large or small equilibrium constant for this reaction? **small, because  $\text{HSO}_4^-$  is a weak acid.**
- Would you define  $\text{HSO}_4^-$  as a proton donor or a proton acceptor? **proton donor, because it is losing a  $\text{H}^+$  ion.**
- In your written equation, is  $\text{HSO}_4^-$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of  $\text{HSO}_4^-$ :  **$\text{SO}_4^{2-}$**
- Identify  $\text{HSO}_4^-$  as polyprotic, amphoteric, or neither. **amphoteric, because it has a negative charge (and can take a proton) and it has hydrogens (and can lose one too).**

# 19 3 Acids And Bases Answer Key Chemistry

**Joseph Clayton Martin**



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## **Table of Contents 19 3 Acids And Bases Answer Key Chemistry**

1. Understanding the eBook 19 3 Acids And Bases Answer Key Chemistry
  - The Rise of Digital Reading 19 3 Acids And Bases Answer Key Chemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying 19 3 Acids And Bases Answer Key Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 19 3 Acids And Bases Answer Key Chemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from 19 3 Acids And Bases Answer Key Chemistry

- Personalized Recommendations
  - 19 3 Acids And Bases Answer Key Chemistry User Reviews and Ratings
  - 19 3 Acids And Bases Answer Key Chemistry and Bestseller Lists
5. Accessing 19 3 Acids And Bases Answer Key Chemistry Free and Paid eBooks
    - 19 3 Acids And Bases Answer Key Chemistry Public Domain eBooks
    - 19 3 Acids And Bases Answer Key Chemistry eBook Subscription Services
    - 19 3 Acids And Bases Answer Key Chemistry Budget-Friendly Options
  6. Navigating 19 3 Acids And Bases Answer Key Chemistry eBook Formats
    - ePub, PDF, MOBI, and More
    - 19 3 Acids And Bases Answer Key Chemistry Compatibility with Devices
    - 19 3 Acids And Bases Answer Key Chemistry Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of 19 3 Acids And Bases Answer Key Chemistry
    - Highlighting and Note-Taking 19 3 Acids And Bases Answer Key Chemistry
    - Interactive Elements 19 3 Acids And Bases Answer Key Chemistry
  8. Staying Engaged with 19 3 Acids And Bases Answer Key Chemistry
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers 19 3 Acids And Bases Answer Key Chemistry
  9. Balancing eBooks and Physical Books 19 3 Acids And Bases Answer Key Chemistry
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection 19 3 Acids And Bases Answer Key Chemistry
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine 19 3 Acids And Bases Answer Key Chemistry
    - Setting Reading Goals 19 3 Acids And Bases Answer Key Chemistry
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of 19 3 Acids And Bases Answer Key Chemistry

- Fact-Checking eBook Content of 19 3 Acids And Bases Answer Key Chemistry
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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