

Acids & Bases HOMEWORK

Homework

1. Propose the correct ion search of the strong acid listed below. Propose the incorrect ion search of the strong base listed below.

- a. **HF**: **Hydrogen fluoride**
- b. **HI**: **Hydrogen iodide**
- c. **HNO₃**: **nitric acid**
- d. **HClO₄**: **perchloric acid**
- e. **HBrO₃**: **hypobromous acid**
- f. **HClO₄**: **chloric acid**
- g. **H₂SO₄**: **hydrochloric acid**
- h. **H₂SO₄**: **hydrobromic acid**
- i. **H₂SO₄**: **hydroiodic acid**
- j. **H₂SO₄**: **hydrochloric acid**
- k. **H₂SO₄**: **hydrobromic acid**
- l. **H₂SO₄**: **hydroiodic acid**

2. Off the acids and bases listed below, write the ones that completely dissociate in water.

Acid	Base	Mixture	Neutral	Acidic	Basic	Neutral	Basic
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3. Write the dissociation for **H₂SO₄**.

4. **HClO₄** will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{HClO}_4 + \text{H}_2\text{O} \longrightarrow \text{H}_3\text{O}^+ + \text{ClO}_4^-$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **small** because HClO_4 is a strong acid.

c. Dissolve **CaCl₂** in water. **CaCl₂** is a problem because it is a **strong** electrolyte? **strong** because it dissociates in H_2O completely.

d. In your pH -ion response, is **HClO₄** acting as a **strong**-strong electrolyte? **strong** because it is a **strong** acid.

e. Dissolve the conjugate base of **NaClO₄**: **ClO₄⁻**

f. Dissolve **NaClO₄** in water. **NaClO₄** is a **strong** electrolyte? **strong** because it dissociates in H_2O completely.

5. H^+ will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{H}^+ + \text{H}_2\text{O} \longrightarrow \text{H}_3\text{O}^+$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **small** because H^+ is a **strong** acid.

c. Dissolve **LiClO₄** in water. **LiClO₄** is a problem because it is a **strong** electrolyte? **strong** because it dissociates in H_2O completely.

d. In your pH -ion response, is **HClO₄** acting as a **strong**-strong electrolyte? **strong** because it is a **strong** acid.

e. Dissolve the conjugate acid of **ClO₄⁻**: **HClO₄**

f. Dissolve **HClO₄** in water. **HClO₄** is a **strong** electrolyte? **strong** because it dissociates in H_2O completely.

6. **HClO₄** will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{HClO}_4 + \text{H}_2\text{O} \longrightarrow \text{H}_3\text{O}^+ + \text{ClO}_4^-$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **small** because **HClO₄** is a **strong** acid.

c. Dissolve **CaCl₂** in water. **CaCl₂** is a problem because it is a **strong** electrolyte? **strong** because it dissociates in H_2O completely.

d. In your pH -ion response, is **HClO₄** acting as a **strong**-strong electrolyte? **strong** because it is a **strong** acid.

e. Dissolve the conjugate base of **HClO₄**: **ClO₄⁻**

f. Dissolve **NaClO₄** in water. **NaClO₄** is a **strong** electrolyte? **strong** because it dissociates in H_2O completely.

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Joseph Clayton Martin

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