

B. 9

SKETCHES  
OF  
UPPER CANADA,  
DOMESTIC,  
LOCAL, AND CHARACTERISTIC:  
TO WHICH ARE ADDED,  
PRACTICAL DETAILS  
FOR THE INFORMATION OF  
EMIGRANTS OF EVERY CLASS;  
AND SOME RECOLLECTIONS OF  
THE UNITED STATES OF AMERICA.

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BY JOHN HOWISON, Esq.

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"In vain does the imagination try to roam at large amidst cultivated plains, for every where the habitations of mankind oppose its wish; but in this desert region the soul delights to bury and lose itself amidst boundless forests—it loves to wander, by the light of the stars, on the borders of immense lakes—to hover on the roaring gulf of terrific cataracts—to fall with the mighty mass of waters—to mix and confound itself, as it were, with the wild sublimities of nature."—*Chateaubriand*.

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1821.

# Sketches Of Upper Canada

**G Thomas**

## **Sketches Of Upper Canada:**

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The Enigmatic Realm of **Sketches Of Upper Canada**: Unleashing the Language is Inner Magic

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the boiling point of an organic liquid is the temperature at which its vapour pressure equals the atmospheric pressure over the liquid or it is the temperature at which the vapour and liquid phases are in equilibrium at a given pressure the boiling point is considered as a criterion of purity of a compound and is useful for identification of

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given that the boiling points of ethanol and methanol 78 °C and 64 °C respectively are far off in a significant way this seems like a valid conclusion conclusion thermometer calibrations are necessary in order to collect reliable information as precise calculations are imperative to gathering correct results

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melting point 1 a push the open end of a capillary tube into the powdered acetamide b move the powder to the closed end of the capillary tube by tapping it on the table repeat until the the powdered acetamide occupies 1 2 mm of the capillary tube end c

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melting point can be used as the identity and purity of an organic solid however it is not the same as the boiling point the boiling point is a characteristic unique of pure liquids the boiling point is the temperature at which the vapor pressure of a liquid is equal to the atmospheric pressure

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50 c so in order to convert an experimental measurement taken in calgary which is at higher altitude than sea level and so is at a lower pressure to that reported for sea level higher pressure one needs to add a correction factor since the boiling point at sea level is higher than that at higher altitudes

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introduction in chemistry one may encounter an unknown substance that need to be determined by physical properties such as boiling point melting point density solubility etc in this experiment i am to figure out the boiling point of a mystery substance in order to identify it chemical responsibility this experiment requires the handling of a bunsen burner boiling

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