

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of a given acid or base solution.
2. Calculate pH values and concentrations of acids and bases and use them to calculate the concentration of a given substance in a given solution.
3. Calculate the pH of a buffer solution.

## Vocabulary

- |                   |                             |                    |
|-------------------|-----------------------------|--------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution   |
| 4. autoionization | 5. acid rain                | 6. buffer solution |
|                   | 7. acid solution            | 8. pH              |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the requested unit or numerical information of the values indicated that are indicated in the given statements. Write complete names for the compounds indicated.

1. What substance is \_\_\_\_\_, a strong hydrogen ion donor?
2. \_\_\_\_\_ is the concentration of ions which produces an acidic solution.
3. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
4. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
5. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
6. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
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18. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
19. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.
20. \_\_\_\_\_ is the concentration of \_\_\_\_\_ in a 0.1 M solution.

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# 19 2 Hydrogen Ions And Acidity Answer Key

**Philip Bovier Hawk, Olaf Bergeim**



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