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web isotopes are those atoms having same atomic number number of protons are same but different mass number number of neutrons differ ions are atoms which contain an overall charge where number of protons number of electrons

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web atoms of the same element i e same atomic number Z that have different numbers of neutrons are called isotopes for example $^{12}_6\text{C}$ of the carbon atoms on earth have 6 neutrons and 6 protons in their nuclei about 1 of the carbon atoms have 7

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we may get an isotope of chlorine with 17 protons and 18 neutrons hence atomic mass 35

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web verified by topper correct option is d atoms having same atomic number but different atomic mass are known as isotopes the difference in the mass is due to the difference in their number of neutrons with the same number of electrons and protons

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web part i answer the questions based on the above reading what is an isotope isotopes are versions of the same element they have the same number of protons and electrons as the element but different mass numbers and number of neutrons what does the number next to isotopes signify the number indicates the isotope's mass number

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element but have unequal numbers of neutrons answer link the nucleus differs as a result of the number of neutrons in the atom the actual structure doesn't differ all that much but its mass does differ thus its density is a little

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