

# Chemistry Study Guide

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1 H	2 He																	18 Ar	19 K	20 Ca											36 Kr	37 Rb	38 Sr											54 Xe	55 Cs	56 Ba															86 Rn																															
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## 🔍 Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles 🧪

## 💡 Matter:

Anything that has **mass** and **occupies space** 🧪🌐

Examples: air, gases, liquids, solids!

## 🔬 Atoms:

Atoms are the **smallest, indivisible particle** 🧪👉

- **Outer part** → *shell/orbit/energy level* → carries the electrons ( $e^-$ )
- **Inner part** → *nucleus* → houses the protons ( $p^+$ ) and neutrons ( $n$ )
- Outer & Inner part → together these are the **subatomic particles**

## 📌 Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	$p^+$	1	+1	Nucleus
Neutron	$n$	1	0	Nucleus
Electron	$e^-$	$\sim 0$	-1	Orbit/Shell

# Section 2 nuclear Chemistry Study Guide Key

**Robert R. Redfield**



## Section 2nuclear Chemistry Study Guide Key:

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