

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa for conjugate acids and bases.
2. Calculate K_a and K_b for polyprotic acids and bases, and vice versa for conjugate acids and bases.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

acid dissociation constant	1. conjugate acid	3. acid ionization constant
acid ionization constant	2. conjugate base	4. base ionization constant

Key Equations

1. $K_a = \frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]}$ (Equation 19.1)
2. $K_b = \frac{[\text{B}][\text{H}_3\text{O}^+]}{[\text{BH}^+]}$ (Equation 19.2)

Part A: Completion

Write the appropriate words in the blanks provided. Each blank may require more than one word. Write complete words, not just parts of words, or symbols.

1. The K_a of a weak acid is 1.5×10^{-4} . The K_b of its conjugate base is _____.
2. The K_b of a weak base is 1.5×10^{-4} . The K_a of its conjugate acid is _____.
3. The K_a of a weak acid is 1.5×10^{-4} . The K_b of its conjugate base is _____.
4. The K_b of a weak base is 1.5×10^{-4} . The K_a of its conjugate acid is _____.
5. The K_a of a weak acid is 1.5×10^{-4} . The K_b of its conjugate base is _____.
6. The K_b of a weak base is 1.5×10^{-4} . The K_a of its conjugate acid is _____.
7. The K_a of a weak acid is 1.5×10^{-4} . The K_b of its conjugate base is _____.
8. The K_b of a weak base is 1.5×10^{-4} . The K_a of its conjugate acid is _____.
9. The K_a of a weak acid is 1.5×10^{-4} . The K_b of its conjugate base is _____.
10. The K_b of a weak base is 1.5×10^{-4} . The K_a of its conjugate acid is _____.

1. K_a of a weak acid is 1.5×10^{-4} .
2. K_b of a weak base is 1.5×10^{-4} .
3. K_a of a weak acid is 1.5×10^{-4} .
4. K_b of a weak base is 1.5×10^{-4} .
5. K_a of a weak acid is 1.5×10^{-4} .
6. K_b of a weak base is 1.5×10^{-4} .
7. K_a of a weak acid is 1.5×10^{-4} .
8. K_b of a weak base is 1.5×10^{-4} .
9. K_a of a weak acid is 1.5×10^{-4} .
10. K_b of a weak base is 1.5×10^{-4} .

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Michael Rajnik



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