

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
3. Calculate the pH of weak acids and bases.

Vocabulary

- acid
- base

- acidic solution
- basic solution

- acidic solution
- basic solution

Key Equations

$$K_a = \frac{[H^+][A^-]}{[HA]} \quad (19.1)$$

$$K_b = \frac{[OH^-][B^+]}{[B]} \quad (19.2)$$

Part A Completion

Write the chemical formula for each of the following acids, bases, and salts. Identify the acid, base, or salt.

1. Hydrochloric acid, HCl , $HClO_4$, HNO_3 , H_2SO_4

2. Acetic acid, CH_3COOH , CH_3COO^- , $CH_3COO^-Na^+$

3. Ammonia, NH_3 , NH_4^+ , $NH_4^+Cl^-$

4. Sodium hydroxide, $NaOH$, $NaOH$, $NaOH$

5. Potassium hydroxide, KOH , KOH , KOH

6. Calcium hydroxide, $Ca(OH)_2$, $Ca(OH)_2$, $Ca(OH)_2$

7. Magnesium hydroxide, $Mg(OH)_2$, $Mg(OH)_2$, $Mg(OH)_2$

8. Barium hydroxide, $Ba(OH)_2$, $Ba(OH)_2$, $Ba(OH)_2$

9. Strontium hydroxide, $Sr(OH)_2$, $Sr(OH)_2$, $Sr(OH)_2$

10. Lithium hydroxide, $LiOH$, $LiOH$, $LiOH$

11. Sodium hydroxide, $NaOH$, $NaOH$, $NaOH$

1. Hydrochloric acid
2. Acetic acid
3. Ammonia
4. Sodium hydroxide
5. Potassium hydroxide
6. Calcium hydroxide
7. Magnesium hydroxide
8. Barium hydroxide
9. Strontium hydroxide
10. Lithium hydroxide
11. Sodium hydroxide

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