

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and pOH and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|---|------------------|
| 1. autoionization | 2. hydronium ion (H_3O^+) | 3. acid solution |
| 4. autoionization | 5. hydronium ion (H_3O^+) | 6. acid solution |
| | | 7. pH |

Key Equation

- | | |
|------------------------------------|--|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the unknown quantity that are indicated in the problem. Be sure to include units in your answers.

1. What substance is _____, a weak hydrogen ion donor?
2. _____ is the concentration of H^+ ions in a solution of 0.001 M HCl .
3. _____ is the concentration of OH^- ions in a solution of 0.001 M NaOH .
4. _____ is the concentration of H^+ ions in a solution of 0.001 M HCl .
5. _____ is the concentration of OH^- ions in a solution of 0.001 M NaOH .
6. _____ is the concentration of H^+ ions in a solution of 0.001 M HCl .
7. _____ is the concentration of OH^- ions in a solution of 0.001 M NaOH .
8. _____ is the concentration of H^+ ions in a solution of 0.001 M HCl .
9. _____ is the concentration of OH^- ions in a solution of 0.001 M NaOH .
10. _____ is the concentration of H^+ ions in a solution of 0.001 M HCl .
11. _____ is the concentration of OH^- ions in a solution of 0.001 M NaOH .

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19 2 Hydrogen Ions And Acidity Worksheet Answers

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