

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa, given K_a or K_b
2. Calculate K_a and K_b for weak acids and bases, and vice versa, given pK_a or pK_b
3. Calculate the percent ionization of weak acids

Vocabulary

- acid dissociation
- acid ionization

- acid ionization constant (K_a)
- acid ionization
- acid ionization

- acid ionization
- acid ionization
- pK_a

Key Equations

- 1. $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1)
- 2. $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2)

- 3. $pK_a = -\log K_a$ (Equation 19.3)
- 4. $pK_b = -\log K_b$ (Equation 19.4)

Part A Completion

Write the appropriate formula or name for each compound, and write its acid-base behavior in water. Indicate whether the compound is a weak acid, weak base, or neutral.

1. CH_3COOH (acetic acid) (Equation 19.1)

2. CH_3COO^- (acetate ion) (Equation 19.1)

3. CH_3COO^- (acetate ion) (Equation 19.1)

4. CH_3COO^- (acetate ion) (Equation 19.1)

5. CH_3COO^- (acetate ion) (Equation 19.1)

6. CH_3COO^- (acetate ion) (Equation 19.1)

7. CH_3COO^- (acetate ion) (Equation 19.1)

8. CH_3COO^- (acetate ion) (Equation 19.1)

9. CH_3COO^- (acetate ion) (Equation 19.1)

10. CH_3COO^- (acetate ion) (Equation 19.1)

11. CH_3COO^- (acetate ion) (Equation 19.1)

12. CH_3COO^- (acetate ion) (Equation 19.1)

13. CH_3COO^- (acetate ion) (Equation 19.1)

14. CH_3COO^- (acetate ion) (Equation 19.1)

15. CH_3COO^- (acetate ion) (Equation 19.1)

16. CH_3COO^- (acetate ion) (Equation 19.1)

17. CH_3COO^- (acetate ion) (Equation 19.1)

18. CH_3COO^- (acetate ion) (Equation 19.1)

19. CH_3COO^- (acetate ion) (Equation 19.1)

20. CH_3COO^- (acetate ion) (Equation 19.1)

21. CH_3COO^- (acetate ion) (Equation 19.1)

22. CH_3COO^- (acetate ion) (Equation 19.1)

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Gary A. Thibodeau, Kevin T. Patton



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