

The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

The total heat capacity of a molecular substance is the sum of each contribution (Figure 16.12 of the text). When equipartition is valid (when the temperature is well above the characteristic temperature of the mode $T \gg \theta_m$) we can estimate the heat capacity by counting the numbers of modes that are active. In gases, all three translational modes are always active and contribute $\frac{3}{2}R$ to the molar heat capacity. If we denote the number of active rotational modes by v_f^* (so for most molecules at normal temperatures $v_f^* = 2$ for linear molecules, and 3 for non-linear molecules), then

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John R. Holum

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 - Advantages of eBooks Over Traditional Books
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 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Physical Chemistry 9th Edition Instructors Solution Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Chemistry 9th Edition Instructors Solution Guide
 - Personalized Recommendations

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- Physical Chemistry 9th Edition Instructors Solution Guide Enhanced eBook Features

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- Interactive Elements Physical Chemistry 9th Edition Instructors Solution Guide

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- Participating in Virtual Book Clubs
- Following Authors and Publishers Physical Chemistry 9th Edition Instructors Solution Guide

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- Benefits of a Digital Library
- Creating a Diverse Reading Collection Physical Chemistry 9th Edition Instructors Solution Guide

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Physical Chemistry 9th Edition Instructors Solution Guide

- Setting Reading Goals Physical Chemistry 9th Edition Instructors Solution Guide
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Physical Chemistry 9th Edition Instructors Solution Guide

- Fact-Checking eBook Content of Physical Chemistry 9th Edition Instructors Solution Guide

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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