

The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

The total heat capacity of a molecular substance is the sum of each contribution (Figure 16.12 of the text). When equipartition is valid (when the temperature is well above the characteristic temperature of the mode $T \gg \theta_m$) we can estimate the heat capacity by counting the numbers of modes that are active. In gases, all three translational modes are always active and contribute $\frac{3}{2}R$ to the molar heat capacity. If we denote the number of active rotational modes by v_f^* (so for most molecules at normal temperatures $v_f^* = 2$ for linear molecules, and 3 for non-linear molecules), then

Physical Chemistry 9th Edition Instructors Solution Guide

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The book delves into Physical Chemistry 9th Edition Instructors Solution Guide. Physical Chemistry 9th Edition Instructors Solution Guide is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Physical Chemistry 9th Edition Instructors Solution Guide, encompassing both the fundamentals and more intricate discussions.

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 - Chapter 5: Conclusion
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