

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and  $pK_a$  and  $pK_b$  for weak acids and bases.
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and  $pK_a$  and  $pK_b$  for weak acids and bases.
3. Calculate the  $pH$  of weak acid and base solutions.

## Vocabulary

- |                 |                                      |                 |
|-----------------|--------------------------------------|-----------------|
| acid            | acid dissociation constant ( $K_a$ ) | acid ionization |
| acidic solution | acidic solution                      | acidic solution |
| acidic solution | acidic solution                      | acidic solution |

## Key Equations

- |   |  |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

## Part A Completion

Write the chemical formula for each of the following acids, bases, and salts. Identify the acid, base, or salt.

1. Hydrochloric acid,  $HCl$ , (Equation 19.1)
2. Acetic acid,  $CH_3COOH$ , (Equation 19.1)
3. Nitric acid,  $HNO_3$ , (Equation 19.1)
4. Sulfuric acid,  $H_2SO_4$ , (Equation 19.1)
5. Phosphoric acid,  $H_3PO_4$ , (Equation 19.1)
6. Hydrofluoric acid,  $HF$ , (Equation 19.1)
7. Hydrobromic acid,  $HBr$ , (Equation 19.1)
8. Hydroiodic acid,  $HI$ , (Equation 19.1)
9. Hydrocyanic acid,  $HCN$ , (Equation 19.1)
10. Hydroperfluoroboric acid,  $HF_2$ , (Equation 19.1)
11. Hydroperfluoroboric acid,  $HF_2$ , (Equation 19.1)
12. Hydroperfluoroboric acid,  $HF_2$ , (Equation 19.1)

1. Hydrochloric acid,  $HCl$
2. Acetic acid,  $CH_3COOH$
3. Nitric acid,  $HNO_3$
4. Sulfuric acid,  $H_2SO_4$
5. Phosphoric acid,  $H_3PO_4$
6. Hydrofluoric acid,  $HF$
7. Hydrobromic acid,  $HBr$
8. Hydroiodic acid,  $HI$
9. Hydrocyanic acid,  $HCN$
10. Hydroperfluoroboric acid,  $HF_2$
11. Hydroperfluoroboric acid,  $HF_2$
12. Hydroperfluoroboric acid,  $HF_2$

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