

19.5

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|-----------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid rain solution |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the unknown quantity that are indicated in the problem. Be sure to include units whenever a unit is requested or implied.

1. What substance is _____, a weak hydrogen ion donor?
2. _____ is the concentration of those species present in water that equal to _____ units.
3. _____ is the concentration of those species present in _____ solutions.
4. _____ is the concentration of those species present in _____ solutions.
5. _____ is the concentration of those species present in _____ solutions.
6. _____ is the concentration of those species present in _____ solutions.
7. _____ is the concentration of those species present in _____ solutions.
8. _____ is the concentration of those species present in _____ solutions.
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17. _____ is the concentration of those species present in _____ solutions.
18. _____ is the concentration of those species present in _____ solutions.
19. _____ is the concentration of those species present in _____ solutions.
20. _____ is the concentration of those species present in _____ solutions.

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