

Multiple Choice:

Selecting the alternative that constitutes the statement or answers the question. What is most commonly observed with water?

1. How many moles of hydrogen gas are in a cylinder at typical conditions, 1000 L?  
 a. 1  
 b. 10  
 c. 100  
 d. 1000

2. At temperature  $T_1$ ,  $1 \text{ mol}$  of  $\text{H}_2\text{O}$  takes twice as long to evaporate as  $1 \text{ mol}$  of  $\text{H}_2$ .  
 a. 100  
 b. 10  
 c. 1  
 d. 0.1

3. The unpaired electron in the lowest ion:  
 a.  $\text{H}^+$  or  $\text{He}^+$  atoms  
 b.  $\text{H}^-$  or  $\text{He}^-$  atoms  
 c.  $\text{H}^+$  or  $\text{He}^+$  ions  
 d.  $\text{H}^-$  or  $\text{He}^-$  ions

4. For temperature  $T_1$ ,  $1 \text{ mol}$  of  $\text{H}_2\text{O}$  takes twice as long to evaporate as  $1 \text{ mol}$  of  $\text{H}_2$ .  
 a. 100  
 b. 10  
 c. 1  
 d. 0.1

5. What is the study of the interaction between a material and particles in motion?  
 a. chemical engineering  
 b. chemical physics  
 c. chemical biology  
 d. chemical geology

6. What is the basic unit of matter?  
 a. a molecule  
 b. an atom  
 c. a cell  
 d. a tissue

7. A substance's empirical formula is  $\text{H}_2\text{O}$ , at the same time, it is also a gas. What is the molecular formula?  
 a.  $\text{H}_2\text{O}$   
 b.  $\text{H}_4\text{O}_2$   
 c.  $\text{H}_2\text{O}_2$   
 d.  $\text{H}_2\text{O}$

8. For temperature  $T_1$ ,  $1 \text{ mol}$  of  $\text{H}_2\text{O}$  takes twice as long to evaporate as  $1 \text{ mol}$  of  $\text{H}_2$ .  
 a. 100  
 b. 10  
 c. 1  
 d. 0.1

9. At temperature  $T_1$ ,  $1 \text{ mol}$  of  $\text{H}_2\text{O}$  takes twice as long to evaporate as  $1 \text{ mol}$  of  $\text{H}_2$ .  
 a. 100  
 b. 10  
 c. 1  
 d. 0.1

10. The molecular formula of water is  $\text{H}_2\text{O}$ . What is the empirical formula?  
 a.  $\text{H}_2\text{O}$   
 b.  $\text{H}_4\text{O}_2$   
 c.  $\text{H}_2\text{O}_2$   
 d.  $\text{H}_2\text{O}$

11. The molecular formula of water is  $\text{H}_2\text{O}$ . What is the empirical formula?  
 a.  $\text{H}_2\text{O}$   
 b.  $\text{H}_4\text{O}_2$   
 c.  $\text{H}_2\text{O}_2$   
 d.  $\text{H}_2\text{O}$

# Multiple Choice Stoichiometry Test

**S Marginson**



## Multiple Choice Stoichiometry Test:

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