

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa, given K_a or K_b
2. Calculate K_a and K_b for weak acids and bases, and vice versa, given pK_a or pK_b
3. Calculate the percent ionization of weak acids

Vocabulary

- acid dissociation
- acid ionization

- acid ionization constant (K_a)
- acid ionization
- acid ionization

- acid ionization
- acid ionization
- pK_a

Key Equations

$$K_a = \frac{[H^+][A^-]}{[HA]} \quad \text{for weak acids}$$

$$K_b = \frac{[OH^-][B^+]}{[B]} \quad \text{for weak bases}$$

Part A Completion

Write the appropriate formula for each of the following acids/bases and give its acid/base ionization constant. Write the formula with appropriate subscripts, charges, and states showing double bonds, where applicable, with appropriate subscripts, charges, and states showing double bonds.

1. Acetic acid, CH_3COOH , $K_a = 1.8 \times 10^{-5}$

2. Hydrofluoric acid, HF , $K_a = 6.6 \times 10^{-4}$

3. Hydrocyanic acid, HCN , $K_a = 4.9 \times 10^{-10}$

4. Hydrobromic acid, HBr , $K_a = 1.0 \times 10^3$

5. Hydrochloric acid, HCl , $K_a = 1.0 \times 10^7$

6. Hydroiodic acid, HI , $K_a = 1.0 \times 10^9$

7. Hydroperfluoroboric acid, HF_2 , $K_a = 1.0 \times 10^5$

8. Hydroperfluoroboric acid, HF_3 , $K_a = 1.0 \times 10^6$

9. Hydroperfluoroboric acid, HF_4 , $K_a = 1.0 \times 10^7$

10. Hydroperfluoroboric acid, HF_5 , $K_a = 1.0 \times 10^8$

11. Hydroperfluoroboric acid, HF_6 , $K_a = 1.0 \times 10^9$

1. Acetic acid, CH_3COOH , $K_a = 1.8 \times 10^{-5}$
2. Hydrofluoric acid, HF , $K_a = 6.6 \times 10^{-4}$
3. Hydrocyanic acid, HCN , $K_a = 4.9 \times 10^{-10}$
4. Hydrobromic acid, HBr , $K_a = 1.0 \times 10^3$
5. Hydrochloric acid, HCl , $K_a = 1.0 \times 10^7$
6. Hydroiodic acid, HI , $K_a = 1.0 \times 10^9$
7. Hydroperfluoroboric acid, HF_2 , $K_a = 1.0 \times 10^5$
8. Hydroperfluoroboric acid, HF_3 , $K_a = 1.0 \times 10^6$
9. Hydroperfluoroboric acid, HF_4 , $K_a = 1.0 \times 10^7$
10. Hydroperfluoroboric acid, HF_5 , $K_a = 1.0 \times 10^8$
11. Hydroperfluoroboric acid, HF_6 , $K_a = 1.0 \times 10^9$

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Table of Contents 19 2 Section Review Hydrogen Ions And Acidity

1. Understanding the eBook 19 2 Section Review Hydrogen Ions And Acidity
 - The Rise of Digital Reading 19 2 Section Review Hydrogen Ions And Acidity
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Section Review Hydrogen Ions And Acidity
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 2 Section Review Hydrogen Ions And Acidity
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Section Review Hydrogen Ions And Acidity
 - Personalized Recommendations
 - 19 2 Section Review Hydrogen Ions And Acidity User Reviews and Ratings

- 19 2 Section Review Hydrogen Ions And Acidity and Bestseller Lists
- 5. Accessing 19 2 Section Review Hydrogen Ions And Acidity Free and Paid eBooks
 - 19 2 Section Review Hydrogen Ions And Acidity Public Domain eBooks
 - 19 2 Section Review Hydrogen Ions And Acidity eBook Subscription Services
 - 19 2 Section Review Hydrogen Ions And Acidity Budget-Friendly Options
- 6. Navigating 19 2 Section Review Hydrogen Ions And Acidity eBook Formats
 - ePub, PDF, MOBI, and More
 - 19 2 Section Review Hydrogen Ions And Acidity Compatibility with Devices
 - 19 2 Section Review Hydrogen Ions And Acidity Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 19 2 Section Review Hydrogen Ions And Acidity
 - Highlighting and Note-Taking 19 2 Section Review Hydrogen Ions And Acidity
 - Interactive Elements 19 2 Section Review Hydrogen Ions And Acidity
- 8. Staying Engaged with 19 2 Section Review Hydrogen Ions And Acidity
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 2 Section Review Hydrogen Ions And Acidity
- 9. Balancing eBooks and Physical Books 19 2 Section Review Hydrogen Ions And Acidity
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 2 Section Review Hydrogen Ions And Acidity
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 19 2 Section Review Hydrogen Ions And Acidity
 - Setting Reading Goals 19 2 Section Review Hydrogen Ions And Acidity
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 19 2 Section Review Hydrogen Ions And Acidity
 - Fact-Checking eBook Content of 19 2 Section Review Hydrogen Ions And Acidity
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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