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contains 2 92 kg of water 1 molality moles solute divided by kilograms solute 0 449 mol kg x 2 92 kg x 1 31108 mol of kbr 2 moles times molar mass equals grams 1 31108 mol 119 0023 g mol 156 g kbr

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web calculate the mass percent mole fraction molality and molarity of hexane c 6 h 14 d 0 660 g cm 3 when it is mixed with heptane c 7 h 12 d 0 684 g cm3 in a 35 0 55 0 ml ratio assume that the volumes are additive

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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web jan 8 2022 mol x 100 the sum of all of the mole percents of a mixture equals 100 mole fraction properties and advantages the mole fraction offers advantages over some of the other units of concentration unlike molarity mole fraction is *worksheet mole fraction ms fleming* - Jun 11 2023

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web converting moles and mass the molecular weight of sodium chloride nacl is 58 44 g mol how many moles of salt are in 13 8 g of sodium chloride express the answer using 3 significant figures stuck use a hint learn for free about math art computer programming economics physics chemistry biology medicine finance history and more

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web a derivatization that achieves this goes as molarity moles of solute volume of solution in liter moles molarity x volume in liter mass moles x molar mass numerical answers 1

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web what is the mole fraction of LiCl in the solution 11 8 practice problem acetone $\text{C}_3\text{H}_6\text{O}$ has a density of 0.791 g/ml and xylene C_8H_{10} has a density of 0.864 g/ml a solution was prepared by mixing acetone and xylene calculate the mole fraction of acetone in a solution if 19.0 ml of acetone was dissolved in 98.4 ml of xylene 7 1

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web since the mole fraction of the O_2 is 0.050 it contributes 0.050 or if you will 5.0 of the total pressure thus total pressure is 0.2105042 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of O_2 are mixed in a flask at room temperature