

Multiple Choice

Selecting the alternative that constitutes the statement or answers the question. What is most commonly observed with water?

1. How many moles of hydrogen gas are in a cylinder at 100°C and 100 atm?
- ☐ 1
☒ 2
☐ 3
☐ 4

2. At 25°C, the K_{sp} of AgCl is 1.8×10^{-10} . What is the molar solubility of AgCl in water?
- ☒ $1.34 \times 10^{-5} \text{ M}$
☐ $1.34 \times 10^{-10} \text{ M}$
☐ $1.34 \times 10^{-5} \text{ M}$
☐ $1.34 \times 10^{-10} \text{ M}$

3. The equilibrium constant for the reaction is:
- ☐ 100 at 25°C
☐ 100 at 100°C
☒ 100 at 25°C
☒ 100 at 100°C

4. The equilibrium constant K_{sp} for AgCl is 1.8×10^{-10} . What is the molar solubility of AgCl in water?
- ☐ 1.34
☐ 1.34
☒ 1.34
☐ 1.34

5. What is the value of the equilibrium constant for the reaction?
- ☐ 100 at 25°C
☒ 100 at 100°C
☐ 100 at 25°C
☐ 100 at 100°C

6. What is the value of the equilibrium constant for the reaction?
- ☐ 1.34
☐ 1.34
☒ 1.34
☐ 1.34

7. The equilibrium constant for the reaction is K_{sp} . What is the value of the equilibrium constant for the reaction?
- ☒ 1.34
☐ 1.34
☐ 1.34
☐ 1.34

8. The equilibrium constant K_{sp} for AgCl is 1.8×10^{-10} . What is the molar solubility of AgCl in water?
- ☒ $1.34 \times 10^{-5} \text{ M}$
☐ $1.34 \times 10^{-10} \text{ M}$
☐ $1.34 \times 10^{-5} \text{ M}$
☐ $1.34 \times 10^{-10} \text{ M}$

9. At 25°C, the K_{sp} of AgCl is 1.8×10^{-10} . What is the molar solubility of AgCl in water?
- ☐ 1.34
☐ 1.34
☒ 1.34
☐ 1.34

10. The equilibrium constant for the reaction is K_{sp} . What is the value of the equilibrium constant for the reaction?
- ☐ 1.34
☒ 1.34
☐ 1.34
☐ 1.34

11. The equilibrium constant K_{sp} for AgCl is 1.8×10^{-10} . What is the molar solubility of AgCl in water?

Multiple Choice Stoichiometry Test

Mary Millhollon, Richard H. Langley



Multiple Choice Stoichiometry Test:

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