

Chemistry Study Guide

Periodic Table of the Elements

<div>1 H</div>	<div>2 He</div>																	<div>18 Ar</div>	<div>19 K</div>	<div>20 Ca</div>											<div>36 Kr</div>	<div>37 Rb</div>	<div>38 Sr</div>											<div>54 Xe</div>	<div>55 Cs</div>	<div>56 Ba</div>															<div>86 Rn</div>																															
<div>3 Li</div>	<div>4 Be</div>																	<div>12 Mg</div>	<div>13 Al</div>	<div>14 Si</div>	<div>15 P</div>	<div>16 S</div>	<div>17 Cl</div>	<div>18 Ar</div>	<div>19 K</div>	<div>20 Ca</div>	<div>21 Sc</div>	<div>22 Ti</div>	<div>23 V</div>	<div>24 Cr</div>	<div>25 Mn</div>	<div>26 Fe</div>	<div>27 Co</div>	<div>28 Ni</div>	<div>29 Cu</div>	<div>30 Zn</div>	<div>31 Ga</div>	<div>32 Ge</div>	<div>33 As</div>	<div>34 Se</div>	<div>35 Br</div>	<div>36 Kr</div>	<div>37 Rb</div>	<div>38 Sr</div>	<div>39 Y</div>	<div>40 Zr</div>	<div>41 Nb</div>	<div>42 Mo</div>	<div>43 Tc</div>	<div>44 Ru</div>	<div>45 Rh</div>	<div>46 Pd</div>	<div>47 Ag</div>	<div>48 Cd</div>	<div>49 In</div>	<div>50 Sn</div>	<div>51 Sb</div>	<div>52 Te</div>	<div>53 I</div>	<div>54 Xe</div>	<div>55 Cs</div>	<div>56 Ba</div>	<div>57 La</div>	<div>58 Ce</div>	<div>59 Pr</div>	<div>60 Nd</div>	<div>61 Pm</div>	<div>62 Sm</div>	<div>63 Eu</div>	<div>64 Gd</div>	<div>65 Tb</div>	<div>66 Dy</div>	<div>67 Ho</div>	<div>68 Er</div>	<div>69 Tm</div>	<div>70 Yb</div>	<div>71 Lu</div>	<div>72 Hf</div>	<div>73 Ta</div>	<div>74 W</div>	<div>75 Re</div>	<div>76 Os</div>	<div>77 Ir</div>	<div>78 Pt</div>	<div>79 Au</div>	<div>80 Hg</div>	<div>81 Tl</div>	<div>82 Pb</div>	<div>83 Bi</div>	<div>84 Po</div>	<div>85 At</div>	<div>86 Rn</div>
<div>57 La</div>	<div>58 Ce</div>	<div>59 Pr</div>	<div>60 Nd</div>	<div>61 Pm</div>	<div>62 Sm</div>	<div>63 Eu</div>	<div>64 Gd</div>	<div>65 Tb</div>	<div>66 Dy</div>	<div>67 Ho</div>	<div>68 Er</div>	<div>69 Tm</div>	<div>70 Yb</div>	<div>71 Lu</div>	<div>72 Hf</div>	<div>73 Ta</div>	<div>74 W</div>	<div>75 Re</div>	<div>76 Os</div>	<div>77 Ir</div>	<div>78 Pt</div>	<div>79 Au</div>	<div>80 Hg</div>	<div>81 Tl</div>	<div>82 Pb</div>	<div>83 Bi</div>	<div>84 Po</div>	<div>85 At</div>	<div>86 Rn</div>																																																															
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🎯 Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles 🧪

💡 Matter:

Anything that has **mass** and **occupies space** 🧪🌐

Examples: air, gases, liquids, solids!

🔬 Atoms:

Atoms are the **smallest, indivisible particle** 🧪👉

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

📌 Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

Section 2 nuclear Chemistry Study Guide Key

EW Minium



Section 2nuclear Chemistry Study Guide Key:

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