

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. autoionization | 5. weak acid | 6. conjugate base |
| | 7. weak base | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the unknown quantity that are indicated in the spaces provided. Be sure to include correct units and significant figures.

1. What substance is _____, a weak hydrogen ion donor?
2. _____ is the concentration of those species present in water that equal to _____.
3. _____ is the concentration of those species present in _____.
4. _____ is the concentration of those species present in _____.
5. _____ is the concentration of those species present in _____.
6. _____ is the concentration of those species present in _____.
7. _____ is the concentration of those species present in _____.
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17. _____ is the concentration of those species present in _____.
18. _____ is the concentration of those species present in _____.
19. _____ is the concentration of those species present in _____.
20. _____ is the concentration of those species present in _____.

1. _____
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19 2 Hydrogen Ions And Acidity Worksheet Answers

Advisory Council on College Chemistry



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