

# Chemistry Study Guide

# Periodic Table of the Elements

<b>H</b> 1 Hydrogen																	<b>He</b> 2 Helium
<b>Li</b> 3 Lithium	<b>Be</b> 4 Beryllium											<b>B</b> 5 Boron	<b>C</b> 6 Carbon	<b>N</b> 7 Nitrogen	<b>O</b> 8 Oxygen	<b>F</b> 9 Fluorine	<b>Ne</b> 10 Neon
<b>Na</b> 11 Sodium	<b>Mg</b> 12 Magnesium											<b>Al</b> 13 Aluminum	<b>Si</b> 14 Silicon	<b>P</b> 15 Phosphorus	<b>S</b> 16 Sulfur	<b>Cl</b> 17 Chlorine	<b>Ar</b> 18 Argon
<b>K</b> 19 Potassium	<b>Ca</b> 20 Calcium	<b>Sc</b> 21 Scandium	<b>Ti</b> 22 Titanium	<b>V</b> 23 Vanadium	<b>Cr</b> 24 Chromium	<b>Mn</b> 25 Manganese	<b>Fe</b> 26 Iron	<b>Co</b> 27 Cobalt	<b>Ni</b> 28 Nickel	<b>Cu</b> 29 Copper	<b>Zn</b> 30 Zinc	<b>Ga</b> 31 Gallium	<b>Ge</b> 32 Germanium	<b>As</b> 33 Arsenic	<b>Se</b> 34 Selenium	<b>Br</b> 35 Bromine	<b>Kr</b> 36 Krypton
<b>Rb</b> 37 Rubidium	<b>Sr</b> 38 Strontium	<b>Y</b> 39 Yttrium	<b>Zr</b> 40 Zirconium	<b>Nb</b> 41 Niobium	<b>Mo</b> 42 Molybdenum	<b>Tc</b> 43 Technetium	<b>Ru</b> 44 Ruthenium	<b>Rh</b> 45 Rhodium	<b>Pd</b> 46 Palladium	<b>Ag</b> 47 Silver	<b>Cd</b> 48 Cadmium	<b>In</b> 49 Indium	<b>Sn</b> 50 Tin	<b>Sb</b> 51 Antimony	<b>Te</b> 52 Tellurium	<b>I</b> 53 Iodine	<b>Xe</b> 54 Xenon
<b>Cs</b> 55 Cesium	<b>Ba</b> 56 Barium		<b>Hf</b> 72 Hafnium	<b>Ta</b> 73 Tantalum	<b>W</b> 74 Tungsten	<b>Re</b> 75 Rhenium	<b>Os</b> 76 Osmium	<b>Ir</b> 77 Iridium	<b>Pt</b> 78 Platinum	<b>Au</b> 79 Gold	<b>Hg</b> 80 Mercury	<b>Tl</b> 81 Thallium	<b>Pb</b> 82 Lead	<b>Bi</b> 83 Bismuth	<b>Po</b> 84 Polonium	<b>At</b> 85 Astatine	<b>Rn</b> 86 Radon
<b>Fr</b> 87 Francium	<b>Ra</b> 88 Radium		<b>Rf</b> 104 Rutherfordium	<b>Db</b> 105 Dubnium	<b>Sg</b> 106 Seaborgium	<b>Bh</b> 107 Bohrium	<b>Hs</b> 108 Hassium	<b>Mt</b> 109 Meitnerium	<b>Ds</b> 110 Darmstadtium	<b>Rg</b> 111 Roentgenium	<b>Cn</b> 112 Copernicium	<b>Uut</b> 113 Ununtrium	<b>Fl</b> 114 Flerovium	<b>Uup</b> 115 Ununpentium	<b>Lv</b> 116 Livermorium	<b>Uus</b> 117 Ununseptium	<b>Uuo</b> 118 Ununoctium
		<div> <div> Lanthanum 57 La </div> <div> Cerium 58 Ce </div> <div> Praseodymium 59 Pr </div> <div> Neodymium 60 Nd </div> <div> Promethium 61 Pm </div> <div> Samarium 62 Sm </div> <div> Europium 63 Eu </div> <div> Gadolinium 64 Gd </div> <div> Terbium 65 Tb </div> <div> Dysprosium 66 Dy </div> <div> Holmium 67 Ho </div> <div> Erbium 68 Er </div> <div> Thulium 69 Tm </div> <div> Ytterbium 70 Yb </div> <div> Lutetium 71 Lu </div> </div>															
		<div> <div> Actinium 89 Ac </div> <div> Thorium 90 Th </div> <div> Protactinium 91 Pa </div> <div> Uranium 92 U </div> <div> Neptunium 93 Np </div> <div> Plutonium 94 Pu </div> <div> Americium 95 Am </div> <div> Curium 96 Cm </div> <div> Berkelium 97 Bk </div> <div> Californium 98 Cf </div> <div> Einsteinium 99 Es </div> <div> Fermium 100 Fm </div> <div> Mendelevium 101 Md </div> <div> Nobelium 102 No </div> <div> Lawrencium 103 Lr </div> </div>															

## Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

## Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

## Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons ( $e^-$ )
- **Inner part** → *nucleus* → houses the protons ( $p^+$ ) and neutrons ( $n$ )
- Outer & Inner part → together these are the **subatomic particles**

## Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	$p^+$	1	+1	Nucleus
Neutron	$n$	1	0	Nucleus
Electron	$e^-$	$\sim 0$	-1	Orbit/Shell

# Section 2nuclear Chemistry Study Guide Key

**George M. Bodner, Harry L. Pardue**



## Section 2nuclear Chemistry Study Guide Key:

*Chemistry, Student Study Guide* John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that s real chemical understanding Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist This text helps you understand how chemical phenomena are governed by what happens at the molecular level apply critical thinking skills to chemical concepts and problems and master the basic mathematical techniques needed for quantitative reasoning You ll see the world as chemists do and learn to appreciate the chemical processes all around us A Fourth Edition with a lot of new perks Revisions include a new early energy chapter revised coverage of bonding expanded coverage of intermolecular forces and increased coverage of multiple equilibria including polyprotic acids New pedagogy strengthens students critical thinking and problem solving skills Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills concepts equations and terms eGrade Plus provides an integrated suite of teaching and learning resources including a complete online version of the text links between problems and relevant sections in the online text practice quizzes the Visual Tutor Interactive LearningWare problems and lab demos as well as homework management and presentation features for instructors Study Guide, Answer Book to Accompany Organic Chemistry Alan S.

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