

# Scavenger Hunt- Nature Walk

Can you find all 20?



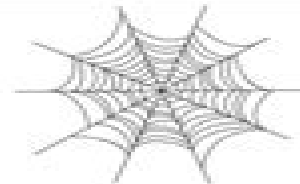
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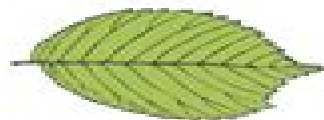
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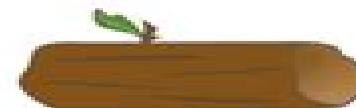
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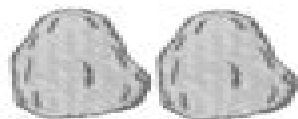
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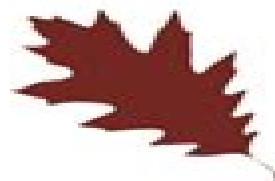
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stick that looks  
like a letter 'y'



flying bug



brown leaf



cloud



dandelion



animal footprints



wild berries



pine cone

# Nature Walk Scavenger Hunt For Kids

**George McDonald, Arthur Frommer**



## **Nature Walk Scavenger Hunt For Kids:**

Fuel your quest for knowledge with is thought-provoking masterpiece, Dive into the World of **Nature Walk Scavenger Hunt For Kids** . This educational ebook, conveniently sized in PDF ( \*), is a gateway to personal growth and intellectual stimulation. Immerse yourself in the enriching content curated to cater to every eager mind. Download now and embark on a learning journey that promises to expand your horizons. .

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web limiting reactant theoretical and percentage yields key  $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$  if a reaction vessel contains 0.1 mol  $\text{KO}_2$  and 0.1 mol  $\text{H}_2\text{O}$ , what is the limiting reactant?

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web limiting reactant: reactant present in an amount lower than required by the reaction stoichiometry, thus limiting the amount of product generated. percent yield: measure of the efficiency of a reaction, expressed as a percentage of the theoretical yield.

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 $\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$

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 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation:  $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}}$

mole  $\text{H}_2\text{S}$  x 18g  $\text{H}_2\text{O}$  i mole  $\text{H}_2\text{O}$  149 4g  $\text{H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137}{149} \times 100 = 91.77\%$  problems 6 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web percent yield is a measure of the efficiency of the experimental design yield efficiency  $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g Al x  $\frac{1 \text{ mol Al}_2\text{O}_3}{2 \text{ mol Al}} \times 101.96 \text{ g/mol} = 81.2 \text{ g}$  c determine the percent yield  $\frac{101.96}{160.0} \times 100 = 63.7\%$

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web 7 Na 2 C 2 O 4 is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H 2 are left over 13

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web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  based on this definition we would expect a percent yield

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how much product will be produced from the limiting reactant calculate how much reactant s remains when the reaction is complete

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