

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa, given  $K_a$  or  $K_b$ .
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa, given  $pK_a$  or  $pK_b$ .
3. Calculate the percent ionization of weak acids.

## Vocabulary

- acid dissociation
- acid ionization

- acid ionization constant ( $K_a$ )
- acid ionization

- base ionization
- base ionization constant ( $K_b$ )

## Key Equations

- 1.  $K_a = \frac{[H^+][A^-]}{[HA]}$  (Equation 19.1)

- 2.  $K_b = \frac{[OH^-][BH^+]}{[B]}$  (Equation 19.2)

## Part A Completion

Write the appropriate formula for each of the following acids/bases and write down its dissociation constant. Write the formula with appropriate subscripts, charges, and states showing dissociation.

1. Acetic acid,  $CH_3COOH$ ,  $K_a = 1.8 \times 10^{-5}$

2. Hydrofluoric acid,  $HF$ ,  $K_a = 6.6 \times 10^{-4}$

3. Hydrocyanic acid,  $HCN$ ,  $K_a = 4.0 \times 10^{-10}$

4. Hydrobromic acid,  $HBr$ ,  $K_a = 1.0 \times 10^3$

5. Hydrochloric acid,  $HCl$ ,  $K_a = 1.0 \times 10^7$

6. Hydroiodic acid,  $HI$ ,  $K_a = 1.0 \times 10^9$

7. Nitric acid,  $HNO_3$ ,  $K_a = 2.4 \times 10^{-4}$

8. Sulfuric acid,  $H_2SO_4$ ,  $K_a = 1.0 \times 10^{-2}$

9. Phosphoric acid,  $H_3PO_4$ ,  $K_a = 7.5 \times 10^{-3}$

10. Acetic acid,  $CH_3COOH$ ,  $K_a = 1.8 \times 10^{-5}$

11. Hydrofluoric acid,  $HF$ ,  $K_a = 6.6 \times 10^{-4}$

1.  $K_a = 1.8 \times 10^{-5}$
2.  $K_a = 6.6 \times 10^{-4}$
3.  $K_a = 4.0 \times 10^{-10}$
4.  $K_a = 1.0 \times 10^3$
5.  $K_a = 1.0 \times 10^7$
6.  $K_a = 1.0 \times 10^9$
7.  $K_a = 2.4 \times 10^{-4}$
8.  $K_a = 1.0 \times 10^{-2}$
9.  $K_a = 7.5 \times 10^{-3}$
10.  $K_a = 1.8 \times 10^{-5}$
11.  $K_a = 6.6 \times 10^{-4}$

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Advances in Inorganic Chemistry and Radiochemistry ,1961-01-01 Advances in Inorganic Chemistry and Radiochemistry

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