

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and pK_a and pK_b for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and pK_a and pK_b for weak acids and bases.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|-----------------|--------------------------------------|-----------------|
| acid | acid dissociation constant (K_a) | acid ionization |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

Part A Completion

Write the chemical formula for each of the following acids, bases, and salts. Identify the acid, base, and salt in each of the following reactions.

1. $HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$

2. $H_2SO_4(aq) + NaOH(aq) \rightarrow Na_2SO_4(aq) + H_2O(l)$

3. $HNO_3(aq) + NaOH(aq) \rightarrow NaNO_3(aq) + H_2O(l)$

4. $H_2CO_3(aq) + NaOH(aq) \rightarrow Na_2CO_3(aq) + H_2O(l)$

5. $H_2SO_4(aq) + NaOH(aq) \rightarrow Na_2SO_4(aq) + H_2O(l)$

6. $HNO_3(aq) + NaOH(aq) \rightarrow NaNO_3(aq) + H_2O(l)$

7. $H_2CO_3(aq) + NaOH(aq) \rightarrow Na_2CO_3(aq) + H_2O(l)$

8. $H_2SO_4(aq) + NaOH(aq) \rightarrow Na_2SO_4(aq) + H_2O(l)$

9. $HNO_3(aq) + NaOH(aq) \rightarrow NaNO_3(aq) + H_2O(l)$

10. $H_2CO_3(aq) + NaOH(aq) \rightarrow Na_2CO_3(aq) + H_2O(l)$

11. $H_2SO_4(aq) + NaOH(aq) \rightarrow Na_2SO_4(aq) + H_2O(l)$

1. $HCl(aq)$
2. $H_2SO_4(aq)$
3. $HNO_3(aq)$
4. $H_2CO_3(aq)$
5. $H_2SO_4(aq)$
6. $HNO_3(aq)$
7. $H_2CO_3(aq)$
8. $H_2SO_4(aq)$
9. $HNO_3(aq)$
10. $H_2CO_3(aq)$
11. $H_2SO_4(aq)$

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