

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on  $K_a$  and  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                   |                             |                    |
|-------------------|-----------------------------|--------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base  |
| 4. autoionization | 5. acid rain                | 6. buffer solution |
| 7. autoionization | 8. acid rain                | 9. pH              |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the equation and/or the numerical value(s) of the constants involved in the reactions described. Use the appropriate symbols.

1. Water autoionization:  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
2. The autoionization of water which produces an  $\text{H}^+$  ion and an  $\text{OH}^-$  ion:  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
3. The autoionization of water which produces an  $\text{H}^+$  ion and an  $\text{OH}^-$  ion:  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
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7. The autoionization of water which produces an  $\text{H}^+$  ion and an  $\text{OH}^-$  ion:  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
8. The autoionization of water which produces an  $\text{H}^+$  ion and an  $\text{OH}^-$  ion:  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
9. The autoionization of water which produces an  $\text{H}^+$  ion and an  $\text{OH}^-$  ion:  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
10. The autoionization of water which produces an  $\text{H}^+$  ion and an  $\text{OH}^-$  ion:  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$

1.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
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