

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HBr~~ **hydrobromic acid**
- ~~HCl~~ **hydrochloric acid**
- ~~HNO₃~~ **nitric acid**
- ~~H₂SO₄~~ **sulfuric acid**
- ~~Strontium hydroxide~~ **LiOH**
- ~~potassium hydroxide~~ **KOH**
- ~~Sodium hydroxide~~ **NaOH**
- ~~calcium hydroxide~~ **Ca(OH)₂**
- ~~ammonium hydroxide~~ **LiOH**

HNO₂ is a WEAK acid & bases will completely dissociate in water

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.

HCl

~~HBr~~

~~HNO₂~~

~~Mg(OH)₂~~

NaOH

NaNO₂

~~NaOH~~

Na₂CO₃

3. What is the formula for hydronium? **H₃O⁺**

4. Cu_2CO_3 will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{Cu}_2\text{CO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{Cu}_2\text{CO}_3^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because Cu_2CO_3 is a weak acid.**
- Would you define Cu_2CO_3 as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is Cu_2CO_3 acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of Cu_2CO_3 : **Cu_2CO_3^-**
- Identify Cu_2CO_3 as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's.**

5. F^- will react with water and gain a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{F}^- + \text{H}_2\text{O} \rightarrow \text{HF} + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F^- is a weak base.**
- Would you define F^- as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H^+ ion.**
- In your written equation, is F^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of F^- : **HF**
- Identify F^- as polyprotic, amphoteric, or neither. **neither**

6. HSO_4^- will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{HSO}_4^- + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{SO}_4^{2-}$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HSO_4^- is a weak acid.**
- Would you define HSO_4^- as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is HSO_4^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of HSO_4^- : **SO_4^{2-}**
- Identify HSO_4^- as polyprotic, amphoteric, or neither. **amphoteric, because it has a negative charge (and will like a proton) and it has H's (and will like to lose one).**

19 3 Acids And Bases Answer Key Chemistry

Henry Watts



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Prentice Hall Chemistry ,2000 *The Chemistry of the Several Natural and Artificial Heterogeneous Compounds Used in Manufacturing Porcelain, Glass, and Pottery* Simeon Shaw,1900 Watts' Dictionary of Chemistry Henry Watts, Henry Forster Morley, Matthew Moncrieff Pattison Muir, 1894 General Chemistry Ralph H. Petrucci, William S. Harwood, 1993 *General Chemistry Principles and Modern Applications* is recognized for its superior problems lucid writing and precision of argument This updated and expanded edition retains the popular and innovative features of previous editions including Feature Problems follow up Integrative and Practice Exercises to accompany every in chapter Example and Focus On application boxes as well as new Keep in Mind marginal notes Topics covered include atoms and the atomic theory chemical compounds and reactions gases Thermochemistry electrons in atoms chemical bonding liquids solids and intermolecular forces chemical kinetics principles of chemical equilibrium acids and bases electrochemistry representative and transitional elements and nuclear and organic chemistry For individuals interested in a broad overview of chemical principles and applications *A Dictionary of Applied Chemistry* Thomas Edward Thorpe, 1894 **A Dictionary of Chemistry and the Allied Branches of Other Sciences** Henry Watts, 1875 *A Dictionary of chemistry and the allied branches of other sciences v. 3, 1882* Henry Watts, 1882 **Chemistry** James Vincent Quagliano, L. M. Vallarino, 1969 *A dictionary of chemistry and the allied branches of other sciences* Henry Watts (F.C.S.), 1882 *Chemistry, Inorganic and Organic* Charles Loudon Bloxam, 1883 *Journal of the Chemical Society* Chemical Society (Great Britain), 1890 Titles of chemical papers in British and foreign journals included in Quarterly journal v 1 12 *Student Solutions Manual for Whitten, Davis, Peck, and Stanley's General Chemistry, 7th Ed* Wendy L. Keeney-Kennicutt, 2004 Contains answers and solutions to all even numbered end of chapter exercises Solutions are divided by section for easy reference by students

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