

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|-----------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid-base reaction |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) for the reaction(s) described. Show the calculation for pH or pOH. Write the chemical formula for the acid or base.

1. What substance is H^+ in water? Hydrogen ion.
2. What substance is OH^- in water? Hydroxide ion.
3. What substance is H^+ in acid? Hydrogen ion.
4. What substance is OH^- in base? Hydroxide ion.
5. What substance is H^+ in acid? Hydrogen ion.
6. What substance is OH^- in base? Hydroxide ion.
7. What substance is H^+ in acid? Hydrogen ion.
8. What substance is OH^- in base? Hydroxide ion.
9. What substance is H^+ in acid? Hydrogen ion.
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97. What substance is H^+ in acid? Hydrogen ion.
98. What substance is OH^- in base? Hydroxide ion.
99. What substance is H^+ in acid? Hydrogen ion.
100. What substance is OH^- in base? Hydroxide ion.

1. H^+ in acid2. OH^- in base3. H^+ in acid4. OH^- in base5. H^+ in acid6. OH^- in base7. H^+ in acid8. OH^- in base9. H^+ in acid10. OH^- in base11. H^+ in acid12. OH^- in base13. H^+ in acid14. OH^- in base15. H^+ in acid16. OH^- in base17. H^+ in acid18. OH^- in base19. H^+ in acid20. OH^- in base

19 2 Hydrogen Ions And Acidity Work Sheet Answers

M Tight



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